

Practicing the Mole -- Even Problems Answer Key

Calculate the mass in grams of each of the following:

2. 8.00 moles of aluminum

$$\frac{8.00 \text{ mol}}{1 \text{ mol}} \left| \begin{array}{c} 26.98 \text{ g} \\ \hline \end{array} \right| = 216 \text{ g Al}$$

6. 7.00 moles of iodine (I_2)

$$\frac{7.00 \text{ mol}}{1 \text{ mol}} \left| \begin{array}{c} 253.8 \text{ g} \\ \hline \end{array} \right| = 1,780 \text{ g } I_2$$

4. 2.00×10^2 moles of chlorine (Cl_2)

$$\frac{200. \text{ mol}}{1 \text{ mol}} \left| \begin{array}{c} 70.90 \text{ g} \\ \hline \end{array} \right| = 14,200 \text{ g } Cl_2$$

8. 9.20 moles of iron

$$\frac{9.20 \text{ mol}}{1 \text{ mol}} \left| \begin{array}{c} 55.85 \text{ g} \\ \hline \end{array} \right| = 514 \text{ g Fe}$$

Calculate the number of moles of atoms in each of the following:

10. 281 g of iron

$$\frac{281 \text{ g}}{55.85 \text{ g}} \left| \begin{array}{c} 1 \text{ mol} \\ \hline \end{array} \right| = 5.03 \text{ mol Fe}$$

14. 93.0 g of phosphorus

$$\frac{93.0 \text{ g}}{30.97 \text{ g}} \left| \begin{array}{c} 1 \text{ mol} \\ \hline \end{array} \right| = 3.00 \text{ mol P}$$

12. 72.0 g of magnesium

$$\frac{72.0 \text{ g}}{24.31 \text{ g}} \left| \begin{array}{c} 1 \text{ mol} \\ \hline \end{array} \right| = 2.96 \text{ mol Mg}$$

16. 1035 g of lead

$$\frac{1035 \text{ g}}{207.2 \text{ g}} \left| \begin{array}{c} 1 \text{ mol} \\ \hline \end{array} \right| = 4.995 \text{ mol Pb}$$

Calculate the number of atoms in each of the following:

18. 5.40 moles of iron

$$\frac{5.40 \text{ mol}}{1 \text{ mole}} \left| \begin{array}{c} 6.022 \times 10^{23} \text{ atoms} \\ \hline \end{array} \right|$$

$$= 3.25 \times 10^{24} \text{ atoms Fe}$$

20. 10.0 moles of titanium

$$\frac{10.0 \text{ mol}}{1 \text{ mole}} \left| \begin{array}{c} 6.022 \times 10^{23} \text{ atoms} \\ \hline \end{array} \right|$$

$$= 6.02 \times 10^{24} \text{ atoms Ti}$$

Answer Key Chemistry Mole Problems

Cathy Duffy

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Table of Contents Answer Key Chemistry Mole Problems

1. Understanding the eBook Answer Key Chemistry Mole Problems
 - The Rise of Digital Reading Answer Key Chemistry Mole Problems
 - Advantages of eBooks Over Traditional Books
2. Identifying Answer Key Chemistry Mole Problems
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer Key Chemistry Mole Problems
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Key Chemistry Mole Problems
 - Personalized Recommendations
 - Answer Key Chemistry Mole Problems User Reviews and Ratings
 - Answer Key Chemistry Mole Problems and Bestseller Lists

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6. Navigating Answer Key Chemistry Mole Problems eBook Formats
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 - Answer Key Chemistry Mole Problems Compatibility with Devices
 - Answer Key Chemistry Mole Problems Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answer Key Chemistry Mole Problems
 - Highlighting and Note-Taking Answer Key Chemistry Mole Problems
 - Interactive Elements Answer Key Chemistry Mole Problems
8. Staying Engaged with Answer Key Chemistry Mole Problems
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answer Key Chemistry Mole Problems
9. Balancing eBooks and Physical Books Answer Key Chemistry Mole Problems
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer Key Chemistry Mole Problems
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Answer Key Chemistry Mole Problems
 - Setting Reading Goals Answer Key Chemistry Mole Problems
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Answer Key Chemistry Mole Problems
 - Fact-Checking eBook Content of Answer Key Chemistry Mole Problems
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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