

**6****Chemical Quantities**

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**A. Completion**

They also encourage students to check prior knowledge of the terms and gain understanding of the concepts introduced in this chapter. Each blank can be completed with a unit, three places, or nothing.

Chemists define units of counting, of mass, and of volume as a single quantity called the 1. This unit measures the "amount of stuff" called 2. One way to measure the amount of a substance is to count the number of 3 of that substance. The amount of representative particles for a sample of a substance is called 4. This number is 5. For example, a sample of nitrogen contains  $6.02 \times 10^{23} 6 of nitrogen.$

Chemists define gram atomic mass as the weight of grams of an element whose mass is equal to the 7 atomic mass. This is the mass of  $6.02 \times 10^{23}$  of carbon, equal to the atomic weight of carbon as 8 g of carbon atoms.

To find the amount in units of a substance represented, we will suppose the 9 of the above-stuffing equals one bag of 10. When the substance the stuff weighs four bags, we divide the 11 of the stuffing by 4 to find the number of bags of stuffing represented by the stuffing weight. The mass of the stuff is called as 12. The formula mass formula mass is 13 units. We would be able to find mass of a sample of stuff without its composition.

The gram formula mass is a mass of a substance composed of a substance as much as the mass of a substance/compound. The definition that measure of mass is a composition of a substance, for example, we can find mass formula mass as 14. For example, mass of the stuff, we can find the mass formula mass as 15 for the substance.

1. mass
2. representative particles
3. Avogadro's Number
4.  $6.02 \times 10^{23}$
5. molecules
6. atomic mass
7. 12
8. gram atomic masses
9. gram molecular mass
10. gram formula mass
11. Molar mass
12. conversion factor
13. chemical formula

# Answer Key Chemical Quantities

**David E. Goldberg**



## Answer Key Chemical Quantities:

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