

Periodic Trends

Can the properties of an element be predicted using a periodic table?

Why?

The periodic table is often considered to be the “best friend” of chemists and chemistry students alike. It includes information about atomic masses and element symbols, but it can also be used to make predictions about atomic size, electronegativity, ionization energies, bonding, solubility, and reactivity. In this activity you will look at a few periodic trends that can help you make those predictions. Like most trends, they are not perfect, but useful just the same.

1. Consider the data in Model 1 on the following page.
 - a. Each element has three numbers listed under it. Which value represents the atomic radius? *The first number shows below the electron shell diagram.*
 - b. What are the units for the atomic radius? *pm—picometers*
 - c. Write a complete sentence to convey your understanding of atomic radius. Hint: You *may* not use the word “radius” in your definition.
Answers will vary. The distance from the center of the atom (nucleus) to the outer electrons.
2. In general, what is the trend in atomic radius as you go down a group in Model 1? Support your answer, using examples from three groups.
The atomic radius gets larger as you go down a group (carbon) in Model 1. Thus, Li is smaller than Be, B is smaller than Al, F is smaller than Cl, etc.
3. Using your knowledge of Coulombic attraction and the structure of the atom, explain the trend in atomic radius that you identified in Question 2. Hint: You should discuss either a change in distance between the nucleus and outer shell of electrons or a change in the number of protons in the nucleus.
Atomic radii increase top to bottom because the number of energy levels increases, increasing the “pull” on electrons from the protons in the nucleus and making the size of the atom larger. The highest occupied energy level has the greatest radius because as the number of electrons increases, more energy levels need to be added further away from the nucleus, making the size of the atom increase.
4. In general, what is the trend in atomic radius as you go across a period (left to right) in Model 1? Support your answer, using examples from two periods.
The atomic radius decreases from left to right across a period. Li is larger than F, Ne is larger than Cl, etc.
5. Using your knowledge of Coulombic attraction and the structure of the atom, explain the trend in atomic radius that you identified in Question 4.
Atomic radii decrease left to right because the number of protons increases, which increases the attractive force on the electrons, and pulls the electrons in the same energy level closer to the nucleus.

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POGIL Shawn R. Simonson, 2023-07-03 Process Oriented Guided Inquiry Learning POGIL is a pedagogy that is based on research on how people learn and has been shown to lead to better student outcomes in many contexts and in a variety of academic disciplines. Beyond facilitating students' mastery of a discipline, it promotes vital educational outcomes such as communication skills and critical thinking. Its active international community of practitioners provides accessible educational development and support for anyone developing related courses. Having started as a process developed by a group of chemistry professors focused on helping their students better grasp the concepts of general chemistry, The POGIL Project has grown into a dynamic organization of committed instructors who help each other transform classrooms and improve student success. They develop curricular materials to assist this process, conduct research expanding what is known about learning and teaching, and provide professional development and collegiality from elementary teachers to college professors. As a pedagogy, it has been shown to be effective in a variety of content areas and at different educational levels. This is an introduction to the process and the community. Every POGIL classroom is different and is a reflection of the uniqueness of the particular context: the institution, department, physical space, student body, and instructor, but follows a common structure in which students work cooperatively in self-managed small groups of three or four. The group work is focused on activities that are carefully designed and scaffolded to enable students to develop important concepts or to deepen and refine their understanding of those ideas or concepts for themselves, based entirely on data provided in class, not on prior reading of the textbook or other introduction to the topic. The learning environment is structured to support the development of process skills such as teamwork, effective communication, information processing, problem solving, and critical thinking. The instructor's role is to facilitate the development of student concepts and process skills, not to simply deliver content to the students. The first part of this book introduces the theoretical and philosophical foundations of POGIL pedagogy and summarizes the literature demonstrating its efficacy. The second part of the book focuses on implementing POGIL, covering the formation and effective management of student teams, offering guidance on the selection and writing of POGIL activities, as well as on facilitation, teaching large classes, and assessment. The book concludes with examples of implementation in STEM and non-STEM disciplines, as well as guidance on how to get started. Appendices provide additional resources and information about The POGIL Project.

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