

Empirical and Molecular Formula Worksheet ANSWER KEY

Write the empirical formula for the following compounds.

- 1) C_6H_6 **CH**
- 2) C_6H_{14} **C_6H_{12}**
- 3) WO_3 **WO_3**
- 4) $\text{C}_2\text{H}_4\text{O}_2$ **CH_2O**
- 5) X_2Y_{12} **X_2Y**
- 6) A compound with an empirical formula of C_2OH_4 and a molar mass of 88 grams per mole. What is the molecular formula of this compound? **$\text{C}_2\text{O}_2\text{H}_4$**
- 7) A compound with an empirical formula of $\text{C}_2\text{H}_4\text{O}$ and a molar mass of 136 grams per mole. What is the molecular formula of this compound? **$\text{C}_2\text{H}_4\text{O}_2$**
- 8) A compound with an empirical formula of CFBrO and a molar mass of 253.8 grams per mole. What is the molecular formula of this compound? **$\text{C}_2\text{F}_2\text{Br}_2\text{O}_2$**
- 9) A compound with an empirical formula of $\text{C}_2\text{H}_4\text{N}$ and a molar mass of 46 grams per mole. What is the molecular formula of this compound? **$\text{C}_2\text{H}_4\text{N}$**
- 10) A well-known reagent in analytical chemistry, dimethylglyoxime, has the empirical formula $\text{C}_2\text{H}_4\text{NO}$. If its molar mass is 116.1 g/mol, what is the molecular formula of the compound? **$\text{C}_2\text{H}_4\text{N}_2\text{O}_2$**
- 11) A certain blue solid contains 36.84% N and 63.16% O. What is the empirical formula of this compound? The ratios are $\text{N}_{1.50}\text{O}_{1.50}$. Since 1.50 is not close to a whole number, we multiply both subscripts by 2. The empirical formula is thus N_2O_3 . (The name is dinitrogen trioxide.) **N_2O_3**
- 12) A sample of indium chloride weighing 0.5000 g is found to contain 0.2404 g of chlorine. What is the empirical formula of the indium compound? **InCl_3**
- 13) An unknown compound was found to have a percent composition as follows: 47.0 % potassium, 14.5 % carbon, and 38.5 % oxygen. What is its empirical formula? If the true molar mass of the compound is 166.22 g/mol, what is its molecular formula? **$\text{K}_2\text{C}_2\text{O}_4$**
- 14) Rubbing alcohol was found to contain 60.0 % carbon, 13.4 % hydrogen, and the remaining mass was due to oxygen. What is the empirical formula of rubbing alcohol? **$\text{C}_3\text{H}_8\text{O}$**

Answer Key To Chemistry A Molecular Approach

**Edward J. Wood,Christopher Smith,W.
Roy Pickering**

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