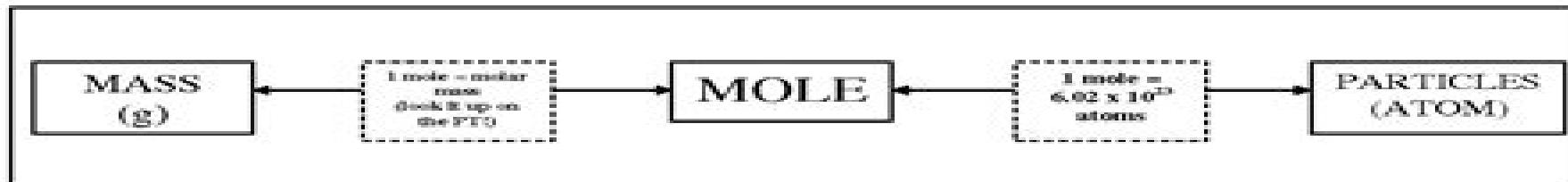


$$1 \text{ mole} = 6.02 \times 10^{23} \text{ atoms} \quad 1 \text{ mole} = \text{atomic mass (g)}$$



### Two Step Problems:

1. What is the mass of  $2.25 \times 10^{23}$  atoms of lead?

$$\frac{2.25 \times 10^{23} \text{ atoms of lead}}{6.02 \times 10^{23} \text{ atoms}} \times \frac{1 \text{ mole}}{1 \text{ mole}} \times \frac{207.2 \text{ g}}{1 \text{ mole}} = 7744.19 \text{ g} = 7740 \text{ g}$$

2. How many atoms are in 10.0g of gold?

$$\frac{10 \text{ g gold}}{197.0 \text{ g}} \times \frac{1 \text{ mole}}{1 \text{ mole}} \times \frac{6.02 \times 10^{23} \text{ atoms}}{1 \text{ mole}} = 3.06 \times 10^{22} \text{ atoms}$$

### PRACTICE PROBLEMS:

(1) How many moles are equal to 60.0g of copper?

$$\frac{60.0 \text{ g Cu}}{63.55 \text{ g Cu}} \times \frac{1 \text{ mole}}{1 \text{ mole}} = 0.95 \text{ mole Cu}$$

(2) How many moles of boron are in a sample containing  $4.25 \times 10^{23}$  atoms of boron?

$$\frac{4.25 \times 10^{23} \text{ atoms of Boron}}{6.02 \times 10^{23} \text{ atoms}} \times \frac{1 \text{ mole}}{1 \text{ mole}} = 0.70 \text{ mole}$$

(3) Convert 2.00 moles of carbon to atoms.

$$2.00 \text{ moles} \times \frac{6.02 \times 10^{23} \text{ atoms}}{1 \text{ mole}} = 1.20 \times 10^{24} \text{ atoms}$$

(4) How many atoms are in 4.0g of potassium?

$$\frac{4.0 \text{ g}}{39.10 \text{ g}} \times \frac{6.02 \times 10^{23} \text{ atoms}}{1 \text{ mole}} = 6.2 \times 10^{22} \text{ atoms}$$

# Answer Key For Moles And Mass

**Frank Schaffer Publications,Joan  
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