

Name _____

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Nuclear Decay

Types of Radiation

1. Complete the table below to compare the properties of alpha, beta, and gamma radiation.

Name	Alpha	Beta	Gamma
Greek letter	α	β^-	γ
Symbol	${}^4_2\text{He}$	${}^0_{-1}\text{e}$	γ
Composition	2 protons, 2 neutrons	an electron	radiation (light)
Charge	0	-1	0
Stopped by	paper, skin, clothing	aluminum	lead

Nuclear Decay Reactions

In problems 2 – 5, fill in the missing particle for each nuclear decay reaction. Then, match each equation to the appropriate type of nuclear decay.



A. alpha emission



B. beta emission



C. electron capture

In problems 6 and 7, write a balanced equation for each nuclear decay reaction.

6. Decay of radium-223 by alpha (α) emission.



7. Decay of iodine-126 by electron (β) emission.



8. How are the atomic number and mass number of a nuclide affected by Beta Decay?

The mass number does not change, but the atomic number increases by 1.

9. How are the atomic number and mass number of a nuclide affected by the emission of gamma radiation?

The atomic number and mass number do not change.

10. Which has a slightly greater mass in a nuclear reaction, the reactants or the products? Why?

It depends on the type of reaction. A fusion reaction will have greater mass in the products because it is combining two nuclei. A fission reaction will not change mass.

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