

15.1**ELECTRON CONFIGURATION IN IONIC BONDING**
SECTION REVIEW**Objectives**

- Use the periodic table to determine the number of valence electrons for an atom and draw its electron-dot structure.
- Describe the bonding and ionic radius of ionic compounds and ionic bonds in terms of electron transfer.

Key Terms

- valence electrons
- electron-dot structure
- ionic bond
- ion

Part A: Completion

After this completion exercise, you will know and understand all the concepts about ionic bonding and ionic structures. Each blank has the completed word or term, where possible, written here.

Elements within the same group of the periodic table have _____ (1) _____.

Similarly, elements in the same period have _____ (2) _____.

_____ (3) _____ is a set of 8 valence electrons called a noble gas.

Electrons in the outermost shell of an atom are called the _____ (4) _____ of the atom.

Electrons in the inner shells are called the _____ (5) _____ of the atom.

Electrons in the outermost shell of an atom are called the _____ (6) _____ of the atom.

Atoms are the smallest units in _____ (7) _____.

Neutral atoms have _____ (8) _____.

Neutral atoms have _____ (9) _____.

Neutral atoms have _____ (10) _____.

Neutral atoms have _____ (11) _____.

Neutral atoms have _____ (12) _____.

Neutral atoms have _____ (13) _____.

Part B: True-False

Classify each of these statements as true or false. If it is false, write the correct answer.

_____ (14) _____ The valence shell gains or loses electrons to become an ion.

_____ (15) _____ The valence shell gains or loses electrons to become a noble gas.

Answer Key Electron Configuration In Ionic Bonding

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