

15.1

ELECTRON CONFIGURATION IN IONIC BONDING

SECTION REVIEW

Objectives

- Use the periodic table to infer the number of valence electrons for an atom and draw its electron dot structure.
- Describe the formation of cations from metals and of anions from nonmetals.

Key Terms

- valence electrons
- electron dot structures
- metal cation
- halide ions

Part A Completion

Use their completion exercises to check your understanding of the concepts and terms discussed and introduced in this section. Each blank can be completed with a term, other phrase, or name here.

Elements within the same group of the periodic table tend to share many features. They have the same number of _____ 1. _____, a number of valence electrons that most indicates how many valence electrons that element has. Gilbert Lewis _____ 2. _____ states that in forming compounds, atoms tend to achieve the electron configuration of a noble gas.

Atoms are held together by _____ 3. _____ of electrical forces. The transfer of valence electrons produces positively charged ions, or _____ 4. _____, and negatively charged ions called _____ 5. _____. The cations of Group 1A elements always have a charge of _____ 6. _____. _____ 7. _____ are produced when atoms of the elements in Group 1A _____ 8. _____ an electron. For transition metals, the _____ 9. _____ of cations may vary.

- _____
- _____
- _____
- _____
- _____
- _____
- _____
- _____
- _____
- _____

Part B True-False

Classify each of these statements as always true, AC: sometimes true, SE: never true, NE:

- _____ 10. The chlorides atom gains seven electrons when it becomes an ion.
- _____ 11. The chemical properties of an element are largely determined by the number of valence electrons that element has.

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