

## 14.3

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate concentrations based on the pH of a given acid or hydrogen ion concentration.
2. Calculate hydrogen ion concentrations from pH and hydroxide ion concentrations from pOH.
3. Calculate the pH of weak acids.

## Vocabulary

- |                   |                               |                       |
|-------------------|-------------------------------|-----------------------|
| 1. autoionization | 2. hydrogen ion concentration | 3. acid solution      |
| 4. autoionization | 5. acid rain                  | 6. acid rain solution |
|                   | 7. acid solution              | 8. pH                 |

## Key Equation

- |                                     |                                     |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

## Part A Completion

Write the hydrogen ion concentration or autoionization constant, if the missing quantity that is indicated in the problem. Express answers in scientific notation with the appropriate number of significant figures.

1. What substance has \_\_\_\_\_ autoionization constant?
2. Autoionization constant of \_\_\_\_\_ is the autoionization of water which produces anhydrous acid equal to \_\_\_\_\_ moles.
3. Hydrochloric acid has a pH of \_\_\_\_\_ or a \_\_\_\_\_ molar solution. \_\_\_\_\_ concentration of \_\_\_\_\_ moles per liter is \_\_\_\_\_ percent by volume.
4. \_\_\_\_\_ autoionization constant is \_\_\_\_\_ moles per liter.
5. \_\_\_\_\_ autoionization constant has a value of \_\_\_\_\_ and \_\_\_\_\_ the product of \_\_\_\_\_ and \_\_\_\_\_ moles.
6. \_\_\_\_\_ autoionization constant is \_\_\_\_\_ moles per liter.

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# 19 2 Hydrogen Ions And Acidity Answer Key

**Valerie C. Scanlon, Tina Sanders**



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