

# Relative formula mass

## Practice worksheet

4.1 Calculating relative formula mass ( $M_r$ )	
Calculate the relative molecular mass of water	
$M_r(\text{H}_2\text{O})$	$= 2 \times 1 + 1 \times 8$ $= 2 \times 1 + 2 \times 16$ $= 2 + 32$ $= 34$
Calculate the $M_r$ of $\text{MgO}$	
$M_r(\text{MgO})$	$= 1 \times \text{Mg} + 1 \times \text{O}$ $= 1 \times 24 + 1 \times 16$ $= 24 + 16$ $= 40$
Calculate the $M_r$ of $\text{CO}_2$	
$M_r(\text{CO}_2)$	$= 1 \times \text{C} + 2 \times \text{O}$ $= 1 \times 12 + 2 \times 16$ $= 12 + 32$ $= 44$
Calculate the $M_r$ of $\text{C}_2\text{H}_6$	
$M_r(\text{C}_2\text{H}_6)$	$= 2 \times \text{C} + 6 \times \text{H}$ $= 2 \times 12 + 6 \times 1$ $= 24 + 6$ $= 30$
Calculate the $M_r$ of $\text{CaCO}_3$	
$M_r(\text{CaCO}_3)$	$= 1 \times \text{Ca} + 1 \times \text{C} + 3 \times \text{O}$ $= 1 \times 40 + 1 \times 12 + 3 \times 16$ $= 40 + 12 + 48$ $= 100$
Calculate the $M_r$ of $\text{H}_2\text{SO}_4$	
$M_r(\text{H}_2\text{SO}_4)$	$= 2 \times \text{H} + 1 \times \text{S} + 4 \times \text{O}$ $= 2 \times 1 + 1 \times 32 + 4 \times 16$ $= 2 + 32 + 64$ $= 98$
Calculate the $M_r$ of $\text{Al}_2\text{O}_3$	
$M_r(\text{Al}_2\text{O}_3)$	$= 2 \times \text{Al} + 3 \times \text{O}$ $= 2 \times 27 + 3 \times 16$ $= 54 + 48$ $= 102$
Calculate the $M_r$ of $\text{K}_2\text{SO}_4$	
$M_r(\text{K}_2\text{SO}_4)$	$= 2 \times \text{K} + 1 \times \text{S} + 4 \times \text{O}$ $= 2 \times 39 + 1 \times 32 + 4 \times 16$ $= 78 + 32 + 64$ $= 174$

4.1 Calculating relative formula mass ( $M_r$ )	
<p>Calculate the <math>M_r</math> of <math>\text{Fe}_2\text{O}_3</math></p> $M_r(\text{Fe}_2\text{O}_3) = 2 \times A_r(\text{Fe}) + 3 \times A_r(\text{O})$ $= 2 \times 56 + 3 \times 16$ $= 112 + 48$ $= 160$	<p>Calculate the <math>M_r</math> of <math>\text{CuSO}_4</math></p> $M_r(\text{CuSO}_4) = 1 \times A_r(\text{Cu}) + 1 \times A_r(\text{S}) + 4 \times A_r(\text{O})$ $= 64 + 32 + 4 \times 16$ $= 64 + 32 + 64$ $= 160$
<p>Calculate the <math>M_r</math> of <math>\text{CaCl}_2</math></p> $M_r(\text{CaCl}_2) = 1 \times A_r(\text{Ca}) + 2 \times A_r(\text{Cl})$ $= 40 + 2 \times 35.5$ $= 40 + 71$ $= 111$	<p>Calculate the <math>M_r</math> of <math>\text{C}_2\text{H}_6\text{O}</math></p> $M_r(\text{C}_2\text{H}_6\text{O}) = 2 \times A_r(\text{C}) + 6 \times A_r(\text{H}) + 1 \times A_r(\text{O})$ $= 2 \times 12 + 6 \times 1 + 16$ $= 24 + 6 + 16$ $= 46$
<p>Calculate the <math>M_r</math> of <math>\text{MgSO}_4</math></p> $M_r(\text{MgSO}_4) = 1 \times A_r(\text{Mg}) + 1 \times A_r(\text{S}) + 4 \times A_r(\text{O})$ $= 24 + 32 + 4 \times 16$ $= 24 + 32 + 64$ $= 120$	<p>Calculate the <math>M_r</math> of <math>\text{HNO}_3</math></p> $M_r(\text{HNO}_3) = 1 \times A_r(\text{H}) + 1 \times A_r(\text{N}) + 3 \times A_r(\text{O})$ $= 1 + 14 + 3 \times 16$ $= 1 + 14 + 48$ $= 63$
<p>Calculate the <math>M_r</math> of <math>\text{CH}_3\text{COOH}</math> (ethanoic acid)</p> <p>Counting atoms: <math>\text{CH}_3\text{COOH}</math></p> $M_r(\text{CH}_3\text{COOH}) = 2 \times A_r(\text{C}) + 4 \times A_r(\text{H}) + 2 \times A_r(\text{O})$ $= 2 \times 12 + 4 \times 1 + 2 \times 16$ $= 24 + 4 + 32$ $= 60$	
<p>Calculate the <math>M_r</math> of <math>\text{Fe}_2\text{O}_3</math></p> $M_r(\text{Fe}_2\text{O}_3) = 2 \times A_r(\text{Fe}) + 3 \times A_r(\text{O})$ $= 2 \times 56 + 3 \times 16$ $= 112 + 48$ $= 160$	

# GCSE Science

## AQA C4.1

# Activity Relative Mass Answers

**RJ Shavelson**



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