

**Experiment 11****Data and Calculations: The Atomic Spectrum of Hydrogen****A. Calculation of the Energy Levels of the Hydrogen Atom**

Energies are to be calculated from Equation 6 for the 10 lowest energy states.

**Table 11.2**

Quantum number, $n$	Energy, $E_n$ [kJ/mol]	Quantum number, $n$	Energy, $E_n$ [kJ/mol]
_____	_____	_____	_____
_____	_____	_____	_____
_____	_____	_____	_____
_____	_____	_____	_____
_____	_____	_____	_____

**B. Calculation of the Wavelengths of the Lines in the Hydrogen Atomic Spectrum**

In the upper half of each box write  $\Delta E$ , the difference in energy in kJ/mole between  $E_{n_u}$  and  $E_{n_l}$ . In the lower half of the box, write the  $\lambda$  in nm associated with that value of  $\Delta E$ .

**Table 11.3**

$n_u$	6	5	4	3	2	1
$n_l$						
1	1275.4 kJ/mol 93.78 nm	1254.50 kJ/mol 94.97 nm	1230.04 kJ/mol 97.25 nm	1166.25 kJ/mol 102.57 nm	984.03 kJ/mol 121.50 nm	
2	291.57 kJ/mol 410.23 nm	275.53 kJ/mol 434.12 nm	246.01 kJ/mol 486.13 nm	182.23 kJ/mol 556.46 nm		
3	109.34 kJ/mol 1094.8 nm	93.5 kJ/mol 1251.0 nm	63.78 kJ/mol 1515.6 nm			
4	45.50 kJ/mol 2075.70 nm	29.52 kJ/mol 4052.48 nm				
5	16.04 kJ/mol 7458.04 nm					

$$\Delta E = E_{n_u} - E_{n_l}$$

$$\lambda \text{ [nm]} = \frac{1.19627 \times 10^5}{\Delta E}$$

(continued on following page)

# Atomic Spectrum Of Hydrogen Lab Answer Key

**K Morrison**



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