

●———— Atomic Structure & Periodic Table ————●

Complete the following table using the periodic table of elements.

Isotope name	Atomic number	Mass number	Number of protons	Number of neutrons	Number of electrons
Hydrogen – 2					
	2	4			
			4	5	
		7	3		
Chlorine – 35					
	19			20	
		24	12		
			41	52	
		107			47
	16			16	
Uranium – 235					
Scandium – 45		45		24	
			31	40	31
	46	106			
Xenon – 136					

Atomic Structure And The Periodic Table Answers

Morris Hein



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Combined Science: Trilogy 9-1 Student Book (GCSE Science 9-1) Ann Daniels,2021-09-20 Exam Board AQA Level Subject GCSE Combined Science Trilogy First teaching September 2016 First exams June 2018 AQA approved

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Understanding Chemistry Fred M. Dewey,1994 **Chemical Interactions** ,2005 **Introduction to General, Organic, and Biochemistry** Morris Hein,1997 The authors recognize that both science and mathematics may be daunting subjects for many students taking this course With this in mind they have anticipated where students might stumble and have paced and organized this text to help them through Their goal is to make the material interesting and relevant so students understand the basic chemical principles related to their career The authors emphasize problem solving and provide a range of practice exercises As in previous editions the text first presents the basic concepts of general chemistry and then moves

into organic and biochemistry In this edition the first two sections have been revised primarily to improve explanations and include new pedagogical features The biochemistry portion has been thoroughly updated to include coverage of many recent developments and emerging technologies in the field **Science** John Michels (Journalist), 1926 **General Science** Richard H. Moyer, 1986 New Technical Books New York Public Library, 1980 Comparative Inorganic Chemistry Bernard John Moody, 1991 An updated and revised edition intended to provide a coherent foundation to the study of inorganic chemistry New sections include the shape of covalent molecules based on the theory of electrical repulsion and the concept of lattice energy Exercises and solutions are included **General Chemistry** Joseph Albert Babor, 1929

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