

CLASS: XI CHEMISTRY
CHAPTER #02
ATOMIC STRUCTURE
BY: SIR AAMIR HULLIO

1. Cathode rays are deflected by
(a) electric field only (b) magnetic field only (c) electric and magnetic field (d) none of these
2. The e/m value for the particles constituting cathode rays is the same regardless of
(a) the gas present in cathode rays tube (b) the metal of which cathode was made
(c) both of these (d) none of these
3. The charge to mass ratio (e/m) of positive particles
(a) varies with the nature of gas in discharge tube (b) is independent of the gas in discharge tube
(c) is constant (d) none of the above
4. A sub atomic particle which has one unit mass and one unit positive charge is known as
(a) hydrogen atom (b) neutron (c) electron (d) proton
5. Atomic number of an element is equal to the number of _____ in the nucleus of the atom.
(a) neutrons (b) protons (c) both the neutrons and protons (d) electrons
6. The mass number of an atom is equal to the number of _____ in the nucleus of an atom
(a) protons (b) neutrons (c) electrons (d) nucleons
7. The mass number and atomic number of Phosphorus atom are 31 and 15 respectively. The number of neutrons in the nucleus is
(a) 15 (b) 16 (c) 31 (d) 46
9. In a sodium atom (atomic number = 11 and mass number = 23), the number of neutrons is
(a) equal to the number of protons (b) less than the number of protons
(c) greater than the number of protons (d) none of these
10. Which of the following is not correct for electromagnetic waves?
(a) the wavelength is the distance between two successive crests
(b) the frequency is the number of waves which pass a given point in one second
(c) the velocity of a wave is the distance covered by the particular wave in one second
(d) all electromagnetic waves have equal wavelengths
11. If Z is the number of proton and A the number of nucleons, then the number of neutrons in an atom is given by
(a) $A + Z$ (b) $A - Z$ (c) $Z - A$ (d) none of these
12. The mass number and atomic number of Phosphorus atom are 31 and 15 respectively. The number of neutrons in the nucleus is
(a) 15 (b) 16 (c) 31 (d) 46
11. Which of the following relations is not correct?
(a) $c = \lambda \times \mu$ (b) $\mu = 1/f$ (c) $\mu = 1/\lambda$ (d) $\lambda = c + \mu$
12. The unit in which wave number is measured
(a) hertz (b) sec^{-1} (c) nanometer (d) cm^{-1}
13. In the spectrum of hydrogen atom, the series which falls in ultraviolet region is
(a) Lyman series (b) Balmer series (c) Paschen series (d) Brackett series
14. The Balmer series in the spectrum of hydrogen atom falls in
(a) ultraviolet region (b) visible region (c) infrared region (d) none of these
21. The radius of first orbit in hydrogen atom is 0.529 \AA . The radius of second orbit is given by
(a) $\frac{1}{2} \times 0.529 \text{ \AA}$ (b) $2 \times 0.529 \text{ \AA}$ (c) $4 \times 0.529 \text{ \AA}$ (d) $8 \times 0.529 \text{ \AA}$

Atomic Structure Ch 4 Packet Answers

**T. W. Graham Solomons, Craig B.
Fryhle, Scott A. Snyder**



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Raymond, 2005-10-07 Finally readers have a shorter less intimidating introduction to general organic and biological chemistry Not only is Raymond s text concise it also takes an integrated approach to presenting important topics in a way that makes the material easier to understand In this approach similarities can be exploited and concepts reinforced The result is that readers see the strong connections that exist between these three branches of chemistry **Study Guide/Selected Solutions Manual** Julie R. Frentrup, 2002-06 Contains a brief overview of every chapter review of skills self tests and the answers and detailed solutions to all odd numbered end of chapter problems in the text book Study Guide to Accompany Basics for Chemistry Martha Mackin, 1982-01-01 Study Guide to Accompany Basics for Chemistry is an 18 chapter text designed to be used with Basics for Chemistry textbook Each chapter contains Overview Topical Outline Skills and Common Mistakes which are all keyed to the textbook for easy cross reference The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms a summary of general concepts and a list of numerical exercises while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook The Fill in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students

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Study Guide, Chemical Principles, Fifth Edition, Zumdahl Paul B. Kelter, Steven S. Zumdahl, 2004-04

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