

## Section Review

## Objectives

- Describe how the amounts of reactants and products change in a chemical system at equilibrium.
- Identify three stresses that can change the equilibrium position of a chemical system.
- Explain what the value of  $K_{\text{eq}}$  indicates about the position of equilibrium.

## Vocabulary

- reversible reaction
- chemical equilibrium
- equilibrium position
- Le Chatelier's principle
- equilibrium constant ( $K_{\text{eq}}$ )

## Key Equation

$$K_{\text{eq}} = \frac{[\text{C}]^c \times [\text{D}]^d}{[\text{A}]^a \times [\text{B}]^b}$$

When  $a\text{A} + b\text{B} \rightleftharpoons c\text{C} + d\text{D}$

## Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short sentence, or number.

## 18 2 Reversible Reactions And Equilibrium Answer Key

**John Crank**

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