

15-2 Practice Problems

What is the molarity of the solution prepared when 14.2 g of potassium dichromate ($K_2Cr_2O_7$) is dissolved in sufficient water to prepare 2.00 L of solution?

$$\frac{14.2 \text{ g } K_2Cr_2O_7}{294.18 \text{ g/mol}} \times \frac{1 \text{ mol}}{2.00 \text{ L}} = 0.0241 \text{ M}$$

How many grams of potassium dichromate ($K_2Cr_2O_7$) are needed to prepare 2.00 L of a 0.100 M solution of potassium dichromate water?

$$2.00 \text{ L} \times 0.100 \text{ M} \times 294.18 \text{ g/mol} = 588 \text{ g}$$

What is the molarity of the solution prepared when 14.2 g of $K_2Cr_2O_7$ and 14.2 g of $K_2Cr_2O_7$ is dissolved in sufficient water to prepare 2.00 L of solution?

$$\frac{14.2 \text{ g } K_2Cr_2O_7}{294.18 \text{ g/mol}} + \frac{14.2 \text{ g } K_2Cr_2O_7}{294.18 \text{ g/mol}} \times \frac{1 \text{ mol}}{2.00 \text{ L}} = 0.0482 \text{ M}$$

25 grams $K_2Cr_2O_7$ and 25 g of solution of potassium dichromate ($K_2Cr_2O_7$) have equal grams of potassium dichromate water for solution

$$25 \text{ g } K_2Cr_2O_7 \times \frac{1 \text{ mol}}{294.18 \text{ g/mol}} \times \frac{1 \text{ mol}}{2.00 \text{ L}} = 0.0425 \text{ M}$$

25 g of $K_2Cr_2O_7$ and potassium water 250 g are dissolved in sufficient water to make 2.00 L of solution, what is the molarity of the solution?

$$\frac{25 \text{ g } K_2Cr_2O_7}{294.18 \text{ g/mol}} \times \frac{1 \text{ mol}}{2.00 \text{ L}} = 0.0425 \text{ M}$$

$$\frac{250 \text{ g } K_2Cr_2O_7}{294.18 \text{ g/mol}} \times \frac{1 \text{ mol}}{2.00 \text{ L}} = 0.425 \text{ M}$$

25 g of $K_2Cr_2O_7$ and 250 g of $K_2Cr_2O_7$ are dissolved in sufficient water to make 2.00 L of solution, what is the molarity of the solution?

$$\frac{25 \text{ g } K_2Cr_2O_7}{294.18 \text{ g/mol}} + \frac{250 \text{ g } K_2Cr_2O_7}{294.18 \text{ g/mol}} \times \frac{1 \text{ mol}}{2.00 \text{ L}} = 0.467 \text{ M}$$

What is the molarity of the solution prepared when 14.2 g of potassium dichromate ($K_2Cr_2O_7$) is dissolved in sufficient water to prepare 2.00 L of solution?

$$\frac{14.2 \text{ g } K_2Cr_2O_7}{294.18 \text{ g/mol}} \times \frac{1 \text{ mol}}{2.00 \text{ L}} = 0.0241 \text{ M}$$

25 grams $K_2Cr_2O_7$ and 25 g of solution of potassium dichromate ($K_2Cr_2O_7$) have equal grams of potassium dichromate water for solution

$$\frac{25 \text{ g } K_2Cr_2O_7}{294.18 \text{ g/mol}} \times \frac{1 \text{ mol}}{2.00 \text{ L}} + \frac{25 \text{ g } K_2Cr_2O_7}{294.18 \text{ g/mol}} \times \frac{1 \text{ mol}}{2.00 \text{ L}} = 0.0425 \text{ M}$$

What is the molarity of a solution containing 14.2 g of $K_2Cr_2O_7$ dissolved in 2.00 L of solution?

$$\frac{14.2 \text{ g } K_2Cr_2O_7}{294.18 \text{ g/mol}} \times \frac{1 \text{ mol}}{2.00 \text{ L}} = 0.0241 \text{ M}$$

What is the molarity of a solution containing 14.2 g of $K_2Cr_2O_7$ dissolved in 2.00 L of solution?

$$\frac{14.2 \text{ g } K_2Cr_2O_7}{294.18 \text{ g/mol}} \times \frac{1 \text{ mol}}{2.00 \text{ L}} = 0.0241 \text{ M}$$

What is the molarity of a solution containing 14.2 g of $K_2Cr_2O_7$ dissolved in 2.00 L of solution?

$$\frac{14.2 \text{ g } K_2Cr_2O_7}{294.18 \text{ g/mol}} \times \frac{1 \text{ mol}}{2.00 \text{ L}} = 0.0241 \text{ M}$$

What is the molarity of a solution containing 14.2 g of $K_2Cr_2O_7$ dissolved in 2.00 L of solution?

$$\frac{14.2 \text{ g } K_2Cr_2O_7}{294.18 \text{ g/mol}} \times \frac{1 \text{ mol}}{2.00 \text{ L}} = 0.0241 \text{ M}$$

What is the molarity of a solution containing 14.2 g of $K_2Cr_2O_7$ dissolved in 2.00 L of solution?

$$\frac{14.2 \text{ g } K_2Cr_2O_7}{294.18 \text{ g/mol}} \times \frac{1 \text{ mol}}{2.00 \text{ L}} = 0.0241 \text{ M}$$

$$\frac{14.2 \text{ g } K_2Cr_2O_7}{294.18 \text{ g/mol}} \times \frac{1 \text{ mol}}{2.00 \text{ L}} = 0.0241 \text{ M}$$

15 2 Practice Problems Answer Key

Christopher Black, Mark Anestis



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