

Review for the Full Year General Chemistry

American Chemical Society Exam

(a) Thermodynamics, (b) Transition Metals, (c) Redox and Electrochemistry, (d) Nuclear and (e) Organic

**Thermodynamics**

84. A 10.0-g sample of silver is heated to 100.0 °C and then added to 20.0 g of water at 23.0 °C in an insulated calorimeter. At thermal equilibrium the temperature of the system was measured as 25.0 °C. What is the specific heat of silver? [specific heat = 4.2 J(g⁻¹ °C⁻¹)]

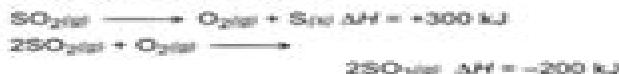
(A) 0.11 J/g K (C) 17 J/g K  
(B) 0.22 J/g K (D) 34 J/g K

85. What is  $\Delta H$  in kJ for the reaction:



(A) -1830 kJ (B) -2284 kJ  
(C) -2510 kJ (D) -1605 kJ

86. Given these equations:



calculate the heat of formation of  $\text{SO}_3\text{(g)}$ .

(A) -500 kJ mol⁻¹ (C) +100 kJ mol⁻¹  
(B) -400 kJ mol⁻¹ (D) +200 kJ mol⁻¹

87. Given:



How much energy is produced in the combustion of 28 g of CO? [Atomic Masses: C 12.0 g mol⁻¹; O 16.0 g mol⁻¹]

(A) 88 kJ (B) 109 kJ  
(C) 175 kJ (D) 284 kJ

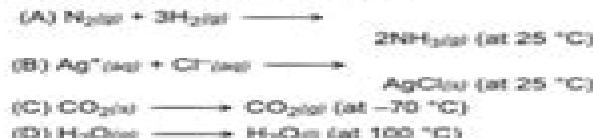
88. Calculate the value of  $\Delta H$  (in  $\text{kJ mol}^{-1}$ ) for the reaction:



Bond Energies (kJ mol⁻¹)	
H-H	435
N≡N	946 (in $\text{N}_2$ )
N-H	389

(A) 2340 kJ of heat absorbed  
(B) 213 kJ of heat absorbed  
(C) 2340 kJ of heat evolved  
(D) 83 kJ of heat evolved

89. Which change is likely to be accompanied by the greatest increase in entropy?



90. For which process is the entropy change per mole the largest at constant temperature?



91. In which process is entropy decreased?

(A) dissolving sugar in water  
(B) expanding a gas  
(C) evaporating a liquid  
(D) freezing water

92. When  $\text{Al}_2\text{O}_3\text{(s)}$  is formed from the elements at standard conditions, the values of  $\Delta H^\circ$  and  $\Delta G^\circ$  at 298 K are -1676 kJ mol⁻¹ and -1577 kJ mol⁻¹, respectively. The standard entropy of formation per mole, in joules per degree, will be

(A) -332 (B) -157 (C) -93.3  
(D) -0.0933 (E) +15.7

93. Vaporization of a liquid is an example of a process for which

(A)  $\Delta H$ ,  $\Delta S$ , and  $\Delta G$  are positive at all temperatures.  
(B)  $\Delta H$  and  $\Delta S$  are positive.  
(C)  $\Delta G$  is negative at low temperatures, positive at high temperatures.  
(D)  $\Delta H = \Delta S$

94. A particular chemical reaction has a negative  $\Delta H$  and negative  $\Delta S$ . Which statement is correct?

(A) The reaction is spontaneous at all temperatures.  
(B) The reaction is nonspontaneous at all temperatures.  
(C) The reaction becomes spontaneous as temperature increases.  
(D) The reaction becomes spontaneous as temperature decreases.

# American Chemical Society Exam Study Guide 2015

**Jay Wackerly, Sarah Zingales, Michael  
Wentzel, Gautam Bhattacharyya, Brett  
McCollum**

## American Chemical Society Exam Study Guide 2015:

**Organic Chemistry Education Research into Practice** Jay Wackerly, Sarah Zingales, Michael Wentzel, Gautam Bhattacharyya, Brett McCollum, 2025-03-25 This Research Topic has three main goals 1 provide a platform for instructors of organic chemistry to showcase evidence based methods and educational theories they have utilized in their classrooms 2 build new and strengthen existing connections between educational researchers and practitioners and 3 highlight how people have used chemical education based research in their teaching practice There are places in the literature dedicated for chemical education research CER however there is not a clear avenue for those that have changed their teaching methods based on published CER and report their experiences Creating this article collection will foster collaboration between chemical education researchers and teachers of organic chemistry This opportunity allows these instructors to share evidence based practices experiences challenges and innovative approaches from CER literature and beyond This Research Topic bridges discipline based education research and the scholarship of teaching and learning which will help advance organic chemistry education and improve student outcomes *Chemistry Education and Contributions from History and Philosophy of Science* Mansoor Niaz, 2015-12-23 This book explores the relationship between the content of chemistry education and the history and philosophy of science HPS framework that underlies such education It discusses the need to present an image that reflects how chemistry developed and progresses It proposes that chemistry should be taught the way it is practiced by chemists as a human enterprise at the interface of scientific practice and HPS Finally it sets out to convince teachers to go beyond the traditional classroom practice and explore new teaching strategies The importance of HPS has been recognized for the science curriculum since the middle of the 20th century The need for teaching chemistry within a historical context is not difficult to understand as HPS is not far below the surface in any science classroom A review of the literature shows that the traditional chemistry classroom curricula and textbooks while dealing with concepts such as law theory model explanation hypothesis observation evidence and idealization generally ignore elements of the history and philosophy of science This book proposes that the conceptual understanding of chemistry requires knowledge and understanding of the history and philosophy of science Professor Niaz's book is most welcome coming at a time when there is an urgently felt need to upgrade the teaching of science The book is a huge aid for adding to the usual way presenting science as a series of mere facts also the necessary mandate to show how science is done and how science through its history and philosophy is part of the cultural development of humanity Gerald Holton Mallinckrodt Professor of Physics these are valence bond VB and molecular orbital MO theories He emphasizes the importance of having rival theories and interpretations in science and its advancement He further argues that this VB MO rivalry is still alive and together the two conceptual frames serve as the tool kit for thinking and doing chemistry in creative manners The author surveys chemistry textbooks in the light of the how the books preserve or not the balance between the two theories in describing various

chemical phenomena This Talmudic approach of conceptual tension is a universal characteristic of any branch of evolving wisdom As such Mansoor s book would be of great utility for chemistry teachers to examine how can they become more effective teachers by recognizing the importance of conceptual tension Sason Shaik Saeree K and Louis P Fiedler Chair in Chemistry Director The Lise Meitner Minerva Center for Computational Quantum Chemistry The Hebrew University of Jerusalem ISRAEL

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### **Table of Contents American Chemical Society Exam Study Guide 2015**

1. Understanding the eBook American Chemical Society Exam Study Guide 2015
  - The Rise of Digital Reading American Chemical Society Exam Study Guide 2015
  - Advantages of eBooks Over Traditional Books
2. Identifying American Chemical Society Exam Study Guide 2015
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an American Chemical Society Exam Study Guide 2015
  - User-Friendly Interface
4. Exploring eBook Recommendations from American Chemical Society Exam Study Guide 2015
  - Personalized Recommendations
  - American Chemical Society Exam Study Guide 2015 User Reviews and Ratings
  - American Chemical Society Exam Study Guide 2015 and Bestseller Lists

5. Accessing American Chemical Society Exam Study Guide 2015 Free and Paid eBooks
  - American Chemical Society Exam Study Guide 2015 Public Domain eBooks
  - American Chemical Society Exam Study Guide 2015 eBook Subscription Services
  - American Chemical Society Exam Study Guide 2015 Budget-Friendly Options
6. Navigating American Chemical Society Exam Study Guide 2015 eBook Formats
  - ePUB, PDF, MOBI, and More
  - American Chemical Society Exam Study Guide 2015 Compatibility with Devices
  - American Chemical Society Exam Study Guide 2015 Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of American Chemical Society Exam Study Guide 2015
  - Highlighting and Note-Taking American Chemical Society Exam Study Guide 2015
  - Interactive Elements American Chemical Society Exam Study Guide 2015
8. Staying Engaged with American Chemical Society Exam Study Guide 2015
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers American Chemical Society Exam Study Guide 2015
9. Balancing eBooks and Physical Books American Chemical Society Exam Study Guide 2015
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection American Chemical Society Exam Study Guide 2015
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine American Chemical Society Exam Study Guide 2015
  - Setting Reading Goals American Chemical Society Exam Study Guide 2015
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of American Chemical Society Exam Study Guide 2015
  - Fact-Checking eBook Content of American Chemical Society Exam Study Guide 2015
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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