

4 • Chemical Equations and Stoichiometry

SIX STOICHIOMETRY PROBLEMS

General Stoichiometry

13. Several brands of antacid tablets use aluminum hydroxide to neutralize excess acid.



[Molar masses: 78.01 36.46 133.4 18.02]

What quantity of HCl, in grams, can a tablet with 0.750 g of Al(OH)_3 consume? What quantity of water is produced?

1.05 g HCl

0.520 g H_2O

17. The equation for one of the reactions in the process of reducing iron ore to the metal is



[Molar masses: 159.7 28.01 55.85 44.01]

- (a) What is the maximum mass of iron, in grams, that can be obtained from 454 g (1.00 lb) of iron(III) oxide?

31.8 g Fe

- (b) What mass of CO is required to reduce the iron(III) oxide to iron metal?

239 g CO

Limiting Reactants

23. The reaction of methane and water is one way to prepare hydrogen:



[Molar masses: 16.04 18.02 28.01 2.02]

If you begin with 995 g of CH_4 and 2510 g of water, what is the maximum possible yield of H_2 ?

CH_4 is LR.... 376 g H_2 is formed

Amadeo Creations Stoichiometry Answers

Mark Steyvers



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