

Acid and Base Worksheet - Answers

- 1) Using your knowledge of the Brønsted-Lowry theory of acids and bases, write equations for the following acid-base reactions and indicate each conjugate acid-base pair:
- a) $\text{HNO}_3 + \text{OH}^- \rightarrow \text{H}_2\text{O} + \text{NO}_3^-$
 HNO_3 and NO_3^- make one pair OH^- and H_2O make the other
- b) $\text{CH}_3\text{NH}_2 + \text{H}_2\text{O} \rightarrow \text{CH}_3\text{NH}_3^+ + \text{OH}^-$
 CH_3NH_2 and CH_3NH_3^+ make one pair OH^- and H_2O make the other
- c) $\text{OH}^- + \text{HPO}_4^{2-} \rightarrow \text{H}_2\text{O} + \text{PO}_4^{3-}$
 HPO_4^{2-} and PO_4^{3-} make one pair OH^- and H_2O make the other
- 2) The compound NaOH is a base by all three of the theories we discussed in class. However, each of the three theories describes what a base is in different terms. Use your knowledge of these three theories to describe NaOH as an Arrhenius base, a Brønsted-Lowry base, and a Lewis base.
- NaOH is an Arrhenius base because it creates OH^- ions when placed in water.
 - NaOH is a Brønsted-Lowry base because it accepts H^+ ions from acids.
 - NaOH is a Lewis base because the lone pairs on the hydroxide ion can be donated to other compounds.
- 3) When hydrogen chloride reacts with ammonia, ammonium chloride is formed. Write the equation for this process, and indicate which of the reagents is the Lewis acid and which is the Lewis base.
- $\text{HCl} + \text{NH}_3 \rightarrow \text{NH}_4\text{Cl}$
 NH_3 is a Lewis base because it uses its lone pair electrons to pull a hydrogen atom from hydrochloric acid.
 HCl is a Lewis acid because it accepts electrons from NH_3 when the H is transferred.
- 4) Write an equation for each of the following reactions and indicate what each reaction produces: (For each of the following oxides, indicate if it is an acid or basic anhydride).
- a. Carbon dioxide and water
 $\text{CO}_2 + \text{H}_2\text{O} \rightarrow \text{H}_2\text{CO}_3$
Acid anhydride
- b. Sodium oxide and water
 $\text{Na}_2\text{O} + \text{H}_2\text{O} \rightarrow 2\text{NaOH(aq)}$
Basic anhydride
- c. Dichlorine trioxide and water
 $\text{Cl}_2\text{O}_3 + \text{H}_2\text{O} \rightarrow 2\text{HClO}_3$
Acidic anhydride
- d. Beryllium oxide and water
 $\text{BeO} + \text{H}_2\text{O} \rightarrow \text{Be(OH)}_2$
Basic anhydride

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Table of Contents Acid Base Review With Answers

1. Understanding the eBook Acid Base Review With Answers
 - The Rise of Digital Reading Acid Base Review With Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Acid Base Review With Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Acid Base Review With Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Acid Base Review With Answers
 - Personalized Recommendations
 - Acid Base Review With Answers User Reviews and Ratings
 - Acid Base Review With Answers and Bestseller Lists
5. Accessing Acid Base Review With Answers Free and Paid eBooks
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 - ePub, PDF, MOBI, and More
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 - Acid Base Review With Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Acid Base Review With Answers
 - Highlighting and Note-Taking Acid Base Review With Answers
 - Interactive Elements Acid Base Review With Answers
- 8. Staying Engaged with Acid Base Review With Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Acid Base Review With Answers
- 9. Balancing eBooks and Physical Books Acid Base Review With Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Acid Base Review With Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Acid Base Review With Answers
 - Setting Reading Goals Acid Base Review With Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Acid Base Review With Answers
 - Fact-Checking eBook Content of Acid Base Review With Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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