

# Acid-Base Equilibrium Practice

- What is the molarity of a solution of nitric acid if 0.216g of barium hydroxide is required to neutralize 20.0ml of nitric acid?
- What is the volume of 1.222M sodium hydroxide required to react with:
  - 32.5ml of 0.569M sulfurous acid
  - 5.00g of oxalic acid,  $\text{H}_2\text{C}_2\text{O}_4$ ?
  - 15.0g of concentrated acetic acid?
- What is the volume of 0.885M hydrochloric acid required to react with:
  - 25.00ml of 0.288M aqueous ammonia?
  - 10.0g of sodium hydroxide?
  - 25.0ml of a solution ( $d=0.928\text{g}/\text{cm}^3$ ) containing 10.0% by mass of methylamine,  $\text{CH}_3\text{NH}_2$ ?
- A capsule of vitamin C, a weak acid, is analyzed by titrating it with a 0.425M sodium hydroxide solution. It is found that 6.20ml of base is required to react with a capsule weighing 0.628g. What is the percentage of vitamin C,  $\text{C}_6\text{H}_8\text{O}_6$ , in the capsule?
- An artificial beverage contains 12.0g of tartaric acid,  $\text{H}_2\text{C}_4\text{H}_4\text{O}_6$ , to achieve tartness. It is titrated with a basic solution that has a density of  $1.045\text{g}/\text{cm}^3$  and contains 5% by mass, KOH. What volume of the basic solution is required?
- Lactic acid,  $\text{C}_3\text{H}_5\text{O}_3$ , is the acid in sour milk. A 0.100g sample of pure lactic acid requires 12.95ml of 0.0857M sodium hydroxide for complete reaction. How many moles of hydroxide ions are required to neutralize one mole of lactic acid?
- An aqueous solution contains dissolved  $\text{C}_6\text{H}_5\text{NH}_2\text{Cl}$  and  $\text{C}_6\text{H}_5\text{NH}_2$ . The concentration of  $\text{C}_6\text{H}_5\text{NH}_2$  is 0.50M and a pH of 4.20.  $K_b = 3.8 \times 10^{-10}$ 
  - Calculate the concentration of  $\text{C}_6\text{H}_5\text{NH}_3^+$  in this buffer solution.
  - Calculate the pH after 4.0grams of HCl is added to 1.0L of solution. Assume no volume change.
- A buffer solution is made up of 355ml each of 0.200M  $\text{NaHCO}_3$  and 0.134M  $\text{Na}_2\text{CO}_3$ , calculate:  $K_a = 5.6 \times 10^{-11}$ 
  - The pH of the buffer
  - The pH of the buffer after the addition of 0.0300mol HCl to 0.710L of buffer solution
  - The pH of the buffer after the addition of 0.0300mol KOH to 0.710L of buffer solution
- Calculate the pH of a solution prepared by mixing 100.0ml of 1.20M ethanolamine,  $\text{C}_2\text{H}_5\text{ONH}_2$ , with 50.0ml of 1.0M HCl.  $K_a$  for  $\text{C}_2\text{H}_5\text{ONH}_3^+ = 6.3 \times 10^{-10}$ .
- Calculate the pH of the following solution:
  - 2.0M  $\text{NaC}_2\text{H}_3\text{O}_2$  / 2.0M  $\text{HC}_2\text{H}_3\text{O}_2$
  - The pH of the buffer solution after 50.0ml of 0.15M LiOH has been added to 100.0ml of the buffer solution.

# Acid Base Equilibrium Practice Problems Multiple Choice

**JS Bruner**



**Acid Base Equilibrium Practice Problems Multiple Choice:**

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