

7 • Atomic Structure**CALCULATION PRACTICE — 2****Formulas and Constants**

$c = \lambda \nu$	$E = h\nu$	$E = \frac{hc}{\lambda}$	$E_n = -\frac{Rhc}{n^2}$	$\lambda = \frac{h}{mv}$	$\frac{1}{\lambda} = R \left(\frac{1}{2^2} - \frac{1}{n^2} \right)$
$c = 3.00 \times 10^8 \text{ m/s}$	$h = 6.626 \times 10^{-34} \text{ J}\cdot\text{s}$	$Rhc = 2.18 \times 10^{-18} \text{ J}$	$R = 1.0974 \times 10^7 \text{ m}^{-1}$		

1. Sketch the electron energy levels ($n=1$ through $n=5$) for the hydrogen atom.



2. Calculate the energy of an electron in the $n=2$ energy level of hydrogen. Calculate the energy of an electron in the $n=3$ energy level. What is the difference in energy of these two levels? If a photon of light had this energy, what would its wavelength be?

$$E_2 = -\frac{2.18 \times 10^{-18} \text{ J}}{4} = \boxed{-5.45 \times 10^{-19} \text{ J}}$$

$$E_3 = -\frac{2.18 \times 10^{-18} \text{ J}}{9} = \boxed{-2.42 \times 10^{-19} \text{ J}}$$

$$\Delta = 3.03 \times 10^{-19} \text{ J}$$

$$\lambda = \frac{hc}{E} = \frac{(6.626 \times 10^{-34} \text{ J}\cdot\text{s})(2.998 \times 10^8 \text{ m/s})}{3.03 \times 10^{-19} \text{ J}} = \boxed{6.56 \times 10^{-7} \text{ m}}$$

3. Use the Rydberg equation above to calculate the wavelength of a photon when $n=3$. How does this compare with your answer in question 2?

$$\frac{1}{\lambda} = 1.0974 \times 10^7 \text{ m}^{-1} \left(\frac{1}{2^2} - \frac{1}{3^2} \right)$$

$$\lambda = \boxed{6.56 \times 10^{-7} \text{ m}}$$

Same

4. An electron moves from the $n=5$ to the $n=1$ quantum level and emits a photon with an energy of $2.093 \times 10^{-18} \text{ J}$. How much energy must the atom absorb to move an electron from $n=1$ to $n=5$? What is the wavelength of this energy?

The same amount
 $-2.093 \times 10^{-18} \text{ J}$

$$\lambda = \frac{hc}{E} = \frac{(6.626 \times 10^{-34} \text{ J}\cdot\text{s})(2.998 \times 10^8 \text{ m/s})}{2.093 \times 10^{-18} \text{ J}} = 9.5 \times 10^{-8} \text{ m} = \boxed{9.491 \times 10^{-8} \text{ m}}$$

5. An electron moves with a velocity of $2.5 \times 10^8 \text{ cm/s}$. What is its wavelength? (The mass of an electron is $9.109 \times 10^{-28} \text{ g}$.)

$$\lambda = \frac{h}{mv} = \frac{6.626 \times 10^{-34} \text{ J}\cdot\text{s}}{(9.109 \times 10^{-28} \text{ g} \times \frac{1 \text{ kg}}{10^3 \text{ g}})(2.5 \times 10^8 \frac{\text{cm}}{\text{s}} \times \frac{1 \text{ m}}{100 \text{ cm}})} = \boxed{2.9 \times 10^{-10} \text{ m}}$$

6. Calculate the wavelength (in nanometers) associated with a $1.0 \times 10^2 \text{ g}$ golf ball moving at $30. \text{ m/s}$ (about 67 mph). How fast must the ball travel to have a wavelength of $5.6 \times 10^{-3} \text{ nm}$?

$$\lambda = \frac{h}{mv} = \frac{6.626 \times 10^{-34} \text{ J}\cdot\text{s}}{(1.0 \times 10^2 \text{ g})(\times \frac{1 \text{ kg}}{10^3 \text{ g}})(30 \frac{\text{m}}{\text{s}})} = \boxed{2.2 \times 10^{-34} \text{ m}} = \boxed{2.2 \times 10^{-25} \text{ nm}}$$

$$5.6 \times 10^{-3} \text{ nm} = \frac{6.626 \times 10^{-34} \text{ J}\cdot\text{s}}{(1.0 \text{ kg}) \times v}$$

$$v = \frac{6.626 \times 10^{-34} \text{ J}\cdot\text{s}}{(1.0 \text{ kg})(5.6 \times 10^{-3} \text{ nm})(\frac{1 \text{ m}}{10^9 \text{ nm}})} = \boxed{1.18 \times 10^{-21} \frac{\text{m}}{\text{s}}}$$

7 Atomic Structure Calculation Practice 1 Answers

Dave Gent

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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touché dans son amour propre il laisse l orgueil et la vanité le dominer et décide de la détromper en séduisant rosette une jeune paysanne sœur de lait de camille espérant ainsi rendre sa cousine jalouse lui donnant rendez vous afin qu elle assiste à la scène

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publiée pour la première fois en 1834 on ne badine pas avec l amour est une des œuvres les plus appréciées d alfred de musset notamment connue pour être le reflet de sa relation avec l auteure george sand pur mélange entre comédie et tragédie et entre réalité et fiction cette pièce a connu une première adaptation

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le drame des rêves de la jeunesse conclusion le drame de l orgueil camille et perdican en effet étaient promis l un à l autre depuis longtemps et la rencontre trop habilement combinée par le baron devait préluder à un mariage heureux mais chacun d eux repousse à son tour l heureux projet en se laissant aller à un mouvement de vanité

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celui ci arrive à ses fins mais lors de leur premier baiser rosette amoureuse de perdican les aperçoit et pousse un cri strident rosette se tue et camille mise au courant du stratagème de perdican le quitte le titre sonne comme un avertissement on

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alfred de musset 1861 la pièce se déroule au château du baron camille la nièce du baron sort du couvent elle a 18 ans elle retrouve le fils de son oncle perdican qui a 21 ans et vient d obtenir un doctorat ils ne s étaient pas vus depuis dix ans le baron souhaite les marier perdican et camille s aiment depuis toujours