

Naming Ionic Compounds

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What are the structural units that make up ionic compounds and how are they named?

Why?

When working in chemistry, it is often convenient to write a chemical in symbols. For example we might write down the substance table salt as NaCl . In talking about chemistry however, it is a bit tacky to say "en-ay see-ell" when we want to refer to a substance. Also, in formal writing we should use the name of the compound rather than its symbols. Therefore we need to learn how to say the proper names of ionic substances.

Model 1 – Ion Charges for Selected Elements

1	H^+									
2	Li^+	He^{2+}						N^{3-}	O^{2-}	F^-
3	Na^+	Mg^{2+}	Transition elements					Al^{3+}	P^{3-}	S^{2-}
4	K^+	Ca^{2+}	Fe^{2+} Fe^{3+}	Ni^{2+} Ni^{3+}	Cu^{+} Cu^{2+}	Zn^{2+}				Br^{3-}
5	Rb^+	Sr^{2+}			Ag^{+}	Hg_2^{2+} Hg^{2+}	S^{2-} S^{3-}			I^-
6		Ba^{2+}					Pb^{2+} Pb^{3+}			

← Cations →

→ Anions →

1. Consider the ions of potassium (K^+) and sulfur (S^{2-}). Write the chemical formula for the ionic compound involving these ions, using the simplest ratio(s) of potassium (K^+) and sulfur (S^{2-}). Keep in mind that the sum of the charges in an ionic compound must equal zero.



2. Consider the ions of iron (Fe^{2+}) and sulfur (S^{2-}). Write chemical formulas for all possible ionic compounds involving these ions, using the simplest ratio(s) of iron (Fe^{2+}) and sulfur (S^{2-}). Keep in mind that the sum of the charges in an ionic compound must equal zero.



Model 2 – Ionic Compound Names (Metals that form one ion)



3. Circle the symbol for the metal in each of the compounds in Model 2. If doing electronically, you can just **BOLD** the symbol. Remember your PPT!

Answer Key For Naming Ionic Compounds Pogil

**Bull, Prince Hycy, Patterson, Gerrelyn
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