

Electron Configurations - KEY

Note: The electron configurations in this worksheet assume that lanthanum (La) is the first element in the 4f block and that actinium (Ac) is the first element in the 5f block. If your periodic table doesn't agree with this, your answers for elements near the f-orbitals may be slightly different (one more/less electron for the f-orbital).

- 1) sodium $1s^2 2s^2 2p^6 3s^1$
- 2) iron $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^6$
- 3) bromine $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^5$
- 4) barium $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2$
- 5) neptunium $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 5p^6 7s^2 5f^3$
also might be written (due to the close energy of d and f and similarities of elements versus predictable
memorized config: $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^6 7s^2 6d^1 5f^4$
- 6) cobalt $[Ar] 4s^2 3d^7$
- 7) silver $[Kr] 5s^2 4d^0$
- 8) tellurium $[Kr] 5s^2 4d^{10} 5p^4$
- 9) radium $[Rn] 7s^2$
- 10) lawrencium $[Rn] 7s^2 5f^{14} 6d^1$
also might be written (due to the close energy of d and f and similarities of elements versus predictable
memorized config: $[Rn] 7s^2 6d^1 5f^{14}$
- 11) $1s^2 2s^2 2p^6 3s^2 3p^4$ sulfur
- 12) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1$ rubidium
- 13) $[Kr] 5s^2 4d^{10} 5p^3$ antimony
- 14) $[Xe] 6s^2 4f^{14} 5d^2$ osmium
- 15) $[Rn] 7s^2 5f^{11}$ einsteinium note: no d's, so the p. table referenced has Ac as 5f¹¹
- 16) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 4d^{10} 4p^3$ not valid (take a look at "4d"; must be 3 d not 4f)
- 17) $1s^2 2s^2 2p^6 3s^2 3d^5$ not valid (3p comes after 3s); also no 3 electrons in S
- 18) $[Ra] 7s^2 5f^2$ not valid (radium isn't a noble gas)
- 19) $[Kr] 5s^2 4d^{10} 5p^3$ valid
- 20) $[Xe]$ not valid (an element can't be its own electron configuration)

Answer Key Electron Configuration

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The Software Encyclopedia 2000 Bowker Editorial Staff, 2000-05

Clear and Simple Chemistry James Berger, Jim Berger, 1986

AQA GCSE Chemistry for Combined Science: Trilogy 9-1 Student Book (GCSE Science 9-1) Ann Daniels, 2021-09-20

Exam Board AQA Level Subject GCSE Combined Science Trilogy First teaching September 2016 First exams June 2018 AQA approved

Fundamentals of Environmental Chemistry, Third Edition Stanley E. Manahan, 2009 This text expands its scope to explore the emerging area that is described as sustainability science and technology which includes green chemistry and industrial ecology. It is designed for those who have little or no knowledge of chemistry but who need the basics of chemical science for their course of study or profession.

Chemistry James C. Hill, 2003 This book assists students through the text material with chapter overviews, learning objectives, review of key terms, cumulative chapter review quizzes and self tests. Included are answers to all Student Guide exercises. Chapter summaries are correlated to those in the Instructor's Resource Manual.

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