

Name: ANSWER KEY Class: _____ Date: _____

Unique Properties of Water!

Word Bank:

| | | | |
|--------------|-------------|-------------|-------------------|
| Adhesion | Evaporation | Polar | Surface tension |
| Cohesion | Freezing | Positive | Universal solvent |
| Condensation | Melting | Sublimation | |
| Dissolve | Negative | | |

1. The electrons are not shared equally between the hydrogen and oxygen atoms of water creating a **Polar** molecule.
2. The polarity of water allows it to **dissolve** most substances. Because of this it is referred to as the **universal solvent**.
3. Water molecules stick to other water molecules. This property is called **cohesion**.
4. Hydrogen bonds form between adjacent water molecules because the **positive** charged hydrogen end of one water molecule attracts the **negative** charged oxygen end of another water molecule.
5. Water molecules stick to other materials due to its polar nature. This property is called **adhesion**.
6. Hydrogen bonds hold water molecules closely together which causes water to have high **surface tension**. This is why water tends to clump together to form drops rather than spread out into a thin film.
7. **Condensation** is when water changes from a gas to a liquid.
8. **Sublimation** is when water changes from a solid directly to a gas.
9. **Freezing** is when water changes from a liquid to a solid.
10. **Melting** is when water changes from a solid to a liquid.
11. **Evaporation** is when water changes from a liquid to a gas.
12. Why does ice float?

Water expands as it freezes, so it is LESS DENSE AS A SOLID.

13. What property refers to water molecules resembling magnets? How are these alike?
Polar bonds create positive and negative ends of the molecule. The positive ends are attracted to other negative ends, creating hydrogen bonds between molecules.

Answer Key To Properties Of Water

Robert R. Redfield

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