

Chemistry

Stoichiometry Reflect #2

Learning Target: I can use stoichiometry to calculate amounts.

Name Ms. Hettie

1a. Balance this equation



1b. Determine how many moles of ammonium oxide was used if the student weighed out 55.2 g of it.

$$55.2 \text{ g (NH}_4\text{)}_2\text{O} \left(\frac{1 \text{ mol A}}{52.10 \text{ g}} \right) = 1.06 \text{ moles (NH}_4\text{)}_2\text{O}$$

1c. Determine the amount of moles of ammonium nitride produced in this reaction.

$$1.06 \text{ moles A} \left(\frac{2 \text{ mol B}}{3 \text{ mol A}} \right) = 0.706 \text{ moles (NH}_4\text{)}_3\text{N}$$

1d. Determine how many grams of ammonium nitride produced in this reaction.

$$0.706 \text{ moles} \left(\frac{48.16 \text{ g}}{1 \text{ mol B}} \right) = 48.16 \text{ g (NH}_4\text{)}_3\text{N}$$

2. Using the following chemical reaction: $2 \text{ Hg} + \text{Cu}(\text{NO}_3)_2 \rightarrow 2 \text{ HgNO}_3 + \text{Cu} \xrightarrow{A}$
and, knowing that 61.6 g of copper was produced in this experiment, how many molecules of mercury nitrate was also produced?

$$61.6 \text{ g Cu} \left(\frac{1 \text{ mol A}}{63.55 \text{ g}} \right) \left(\frac{2 \text{ mol B}}{1 \text{ mol A}} \right) \left(\frac{6.02 \times 10^{23} \text{ mc}}{1 \text{ mol B}} \right) = 1.17 \times 10^{24} \text{ mc HgNO}_3$$

3. Challenge...

What is the percent yield of the following reaction if 12.8 g of copper (II) chloride was used in the lab, and 7.75 g of aluminum chloride was measured at the end of the experiment? (hint: you must calculate the theoretical yield of aluminum chloride first!!)



$$12.8 \text{ g A} \left(\frac{1 \text{ mol A}}{134.45 \text{ g}} \right) \left(\frac{2 \text{ mol B}}{3 \text{ mol A}} \right) \left(\frac{133.33 \text{ g}}{1 \text{ mol B}} \right) = 8.46 \text{ g AlCl}_3$$

Theoretical

$$\begin{aligned} \text{yield} &= \frac{7.75 \text{ g}}{8.46 \text{ g}} \\ &= 91.4\% \end{aligned}$$

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Answer Key To Mixed Stoichiometry Practice

LL Leslie

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