

## SOLUBILITY CURVE WORKSHEET KEY

Use your solubility curve graphs provided to answer the following questions.

1. What are the customary units of solubility on solubility curves? Degress Celsius and grams of solute/100g of water
2. Define solubility. A measure of how much solute can dissolve in a given amount of solvent.
3. According to the graph, the solubility of any substance changes as temperature changes.
4. List the substances whose solubility decreases as temperature increases.  $\text{NH}_3$  and  $\text{Ce}_2(\text{SO}_4)_3$
5. Which substance is least affected by temperature changes?  $\text{NaCl}$
6. How many grams of ammonium chloride ( $\text{NH}_4\text{Cl}$ ) at  $50^\circ\text{C}$ ? 50g
7.  $\text{NaCl}$  and  $\text{KClO}_3$  have the same solubility at approximately  $78^\circ\text{C}$ .
8. Which compound is least soluble in water at  $10^\circ\text{C}$ ?  $\text{KClO}_3$
9. How many grams of  $\text{KNO}_3$  can be dissolved at  $50^\circ\text{C}$ ? 80g
10. Are the following solutions unsaturated, saturated, or supersaturated?
  - a. 45g of  $\text{NaNO}_3$  in 100 g of water at  $30^\circ\text{C}$ . saturated
  - b. 60g of  $\text{KClO}_3$  in 100 g of water at  $90^\circ\text{C}$ . supersaturated
11. How many grams of sodium chloride,  $\text{NaCl}$  are required to saturate 100 grams of water at  $100^\circ\text{C}$ ? 40g
12. How many grams of  $\text{NaNO}_3$  are required to saturate 100 grams of water at  $75^\circ\text{C}$ ? 140g
13. How many grams of  $\text{KI}$  will saturate water at  $20^\circ\text{C}$ ? 33g
14. At what temperature would 25g of potassium chlorate ( $\text{KClO}_3$ ) dissolve?  $60^\circ\text{C}$
15. At what temperature would 60g of  $\text{NH}_4\text{Cl}$  dissolve?  $70^\circ\text{C}$
16. 89 g  $\text{NaNO}_3$  is prepared at  $30^\circ\text{C}$ .
  - a) Will all of the salt dissolve? No
  - b) What mass of  $\text{NaNO}_3$  will dissolve at this temperature? 95g
17. If 50 grams of  $\text{NH}_4\text{Cl}$  is dissolved at  $50^\circ\text{C}$ , how many additional grams  $\text{NH}_4\text{Cl}$  would be needed to make the solution saturated at  $80^\circ\text{C}$ ? 15g
18. At  $50^\circ\text{C}$ , how many grams of  $\text{KNO}_3$  will dissolve? 80g
19. At  $70^\circ\text{C}$ , how many grams of cerium (III) sulfate ( $\text{Ce}_2(\text{SO}_4)_3$ ) dissolve? 5g
20. Determine if each of the following is unsaturated, saturated, or supersaturated.
  - a. 55g of  $\text{NH}_3$  at  $20^\circ\text{C}$ . supersaturated
  - b. 10g of  $\text{Ce}_2(\text{SO}_4)_3$  at  $10^\circ\text{C}$ . unsaturated
  - c. 110g of  $\text{KNO}_3$  at  $60^\circ\text{C}$ . supersaturated
  - d. 65g of  $\text{NH}_4\text{Cl}$  at  $80^\circ\text{C}$ . saturated
  - e. 12g of  $\text{NH}_3$  at  $90^\circ\text{C}$ . supersaturated
  - f. 78g of  $\text{NaNO}_3$  at  $10^\circ\text{C}$ . saturated
  - g. 145g of  $\text{NaNO}_3$  at  $80^\circ\text{C}$ . saturated
  - h. 35g of  $\text{NaCl}$  at  $100^\circ\text{C}$ . unsaturated

# Answer Key To Lab Solubility Data Sheet

**J. A. Beran**



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## **Table of Contents Answer Key To Lab Solubility Data Sheet**

1. Understanding the eBook Answer Key To Lab Solubility Data Sheet
  - The Rise of Digital Reading Answer Key To Lab Solubility Data Sheet
  - Advantages of eBooks Over Traditional Books
2. Identifying Answer Key To Lab Solubility Data Sheet
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Answer Key To Lab Solubility Data Sheet
  - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Key To Lab Solubility Data Sheet
  - Personalized Recommendations
  - Answer Key To Lab Solubility Data Sheet User Reviews and Ratings
  - Answer Key To Lab Solubility Data Sheet and Bestseller Lists
5. Accessing Answer Key To Lab Solubility Data Sheet Free and Paid eBooks
  - Answer Key To Lab Solubility Data Sheet Public Domain eBooks
  - Answer Key To Lab Solubility Data Sheet eBook Subscription Services
  - Answer Key To Lab Solubility Data Sheet Budget-Friendly Options
6. Navigating Answer Key To Lab Solubility Data Sheet eBook Formats

- ePub, PDF, MOBI, and More
  - Answer Key To Lab Solubility Data Sheet Compatibility with Devices
  - Answer Key To Lab Solubility Data Sheet Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Answer Key To Lab Solubility Data Sheet
  - Highlighting and Note-Taking Answer Key To Lab Solubility Data Sheet
  - Interactive Elements Answer Key To Lab Solubility Data Sheet
8. Staying Engaged with Answer Key To Lab Solubility Data Sheet
- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Answer Key To Lab Solubility Data Sheet
9. Balancing eBooks and Physical Books Answer Key To Lab Solubility Data Sheet
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection Answer Key To Lab Solubility Data Sheet
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Answer Key To Lab Solubility Data Sheet
- Setting Reading Goals Answer Key To Lab Solubility Data Sheet
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Answer Key To Lab Solubility Data Sheet
- Fact-Checking eBook Content of Answer Key To Lab Solubility Data Sheet
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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