

STOICHIOMETRY: MOLE-MOLE PROBLEMS



How many moles of hydrogen are needed to completely react with two moles of nitrogen?



6 moles H_2



How many moles of oxygen are produced by the decomposition of six moles of potassium chlorate?



9 moles of O_2



How many moles of hydrogen are produced from the reaction of three moles of zinc with an excess of hydrochloric acid?



3 moles of H_2



How many moles of oxygen are necessary to react completely with four moles of propane (C_3H_8)?



20 mol of O_2



How many moles of potassium nitrate are produced when two moles of potassium phosphate react with two moles of aluminum nitrate?



6 moles of KNO_3

Answer Key Stoichiometry Packet

Cosmas C Mbazi

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Calculating in Chemistry Cosmas C Mbuzi, 2022-09-20 This book is intended to help students fully grasp calculations involving reacting quantities Students in higher courses may find it a helpful revision and enhance their clarity The book completely discusses the key topics in basic stoichiometry including the mole concept reacting quantities and empirical and molecular formulas It begins with the basic concepts and formulas required to convert various quantities to moles or amount of substance This is particularly useful to a beginner Chapter 2 describes how to calculate reacting quantities and therefore provides a general step by step framework or approach by which to solve these problems The chapter also describes and applies the concept of a limiting reagent Two methods of determining a limiting reagent are explained and illustrated The concept of limiting reagent is extended to reacting volumes of gases with a short cut method A

short cut method for solving reacting quantities involving masses and volumes of gases is also given Chapter 3 describes the calculations involved in the practical determination of molecular and empirical formulas A clear meaning of percentage composition of mass is provided and used to solve problems in a step by step manner In chapter 4 we discuss percentage yield and purity A number of examples are given to illustrate how formulas of yield and purity are used in various circumstances A student will find these examples helpful in relating different formulas for percentage purity The last chapter introduces a graphical method for reacting quantities The method may provide a new way of looking at chemical reactions Examples are given to illustrate the method including how it can be used to determine limiting reagents It is hoped that the book will provide all the necessary knowledge and skills to students studying an introductory chemistry course Teachers may also find this book a good resource for their lessons in stoichiometry **Stoichiometry** Georg G. Behrens,2016

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