

Periodic Trends

ATOMIC RADIUS

1. What trend in atomic radius do you see as you go down a group/family on the periodic table?
2. What causes this trend?

3. What trend in atomic radius do you see as you go across a period/row on the periodic table?
4. What causes this trend?

5. Circle the atom in each pair that has the largest atomic radius.
a) Al B b) S O c) Br Cl
d) Na Al e) O F f) Mg Ca

6. Put the following elements in order from smallest to largest atomic radius **and explain why**:
C, O, Sn, Sr.

ELECTRONEGATIVITY

7. Define electronegativity

8. How does the ionic radius of a nonmetal compare with its atomic radius?

9. What trend in electronegativity do you see as you go down a group/family on the periodic table?

10. What causes this trend?

11. What trend in electronegativity do you see as you go across a period/row on the periodic table?

12. What causes this trend?

13. Circle the atom in each pair that has the greater electronegativity.
a) Ca Ga b) Li O c) Cl S d) Br As e) Ba Sr f) O S

GENERAL QUESTIONS

14. Which group tends to form +1 ions? _____
15. Which group tends to form +2 ions? _____
16. Which group tends to form -1 ions? _____
17. Which group tends not to form ions or react? _____

Answer Key For Periodic Trends Test

Claire Aldridge, Karl Lee

Answer Key For Periodic Trends Test:

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