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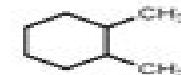
55 multiple-choice questions (3 pts each) + 6 fill-in questions (40 pts) = Total (205 pts).

1) How many unpaired electrons are present in the electron configuration of the O atom?  
A) none      B) one      C) two      D) three      E) four2) Which of the following contain(s) polar covalent bonds?  
A) H<sub>2</sub>      B) NaI      C) Cl<sub>2</sub>      D) H<sub>2</sub>O      E) both A and C3) What is the formal charge and hybridization for the highlighted Carbon?  
A) 0, sp<sup>2</sup>      B) +1, sp<sup>2</sup>      C) -1, sp      D) +1, sp      E) 0, sp       $\text{H}_3\text{C}-\text{N}\equiv\text{C}^+$ 4) What of the following best describes the hybridization of C in carbocation CH<sub>3</sub><sup>+</sup>?  
A) sp<sup>3</sup>      B) sp<sup>2</sup>      C) sp      D) s5) what orbital is the nitrogen lone pair residing in the following molecule?  
A) sp<sup>3</sup>      B) sp<sup>2</sup>      C) sp      D) p      E) s

6) Identify the most acidic carboxylic acid.



Answer is E

7) The conjugate acid of H<sub>2</sub>O is  
A) HO<sup>-</sup>      B) H<sub>3</sub>O      C) H<sub>3</sub>O<sup>+</sup>      D) H<sub>2</sub>O      E) H<sub>2</sub>O<sup>+</sup>8) How many potential stereoisomers exist for the following molecule?  
A) 1      B) 2      C) 3      D) 4      E) 5

9) what is the name of the following compound?

A) *trans*-4-ethylcyclohexan-1-ol      B) *cis*-4-ethylhexan-1-ol  
C) *trans*-4-methylcyclohexan-1-ol      D) *cis*-4-ethylcyclohexan-1-ol  
E) *trans*-4-ethylcyclopentan-1-ol

10) which of the following step is part of the propagation in the methane chlorination?

A)  $\text{Cl}-\text{Cl} \longrightarrow 2 \text{Cl}^{\cdot}$       B)  $\text{H}_3\text{C}^{\cdot} + \text{Cl}-\text{Cl} \longrightarrow \text{H}_3\text{C}-\text{Cl} + \text{Cl}^{\cdot}$   
C)  $\text{H}_3\text{C}-\text{H} + \text{Cl}^{\cdot} \longrightarrow \text{H}_3\text{C}^{\cdot} + \text{H}-\text{Cl}$       D) B and C

Answer is D

11) Based on the Hammond postulate, which of the following transition state best explains the following reaction step?

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**Thomas Bliss Stillman**

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