

Experiment 25

Advance Study Assignment: pH, Its Measurement and Applications

1. A solution of a weak acid was tested with the indicators used in this experiment. The colors observed were as follows:

Methyl violet	violet	Congo red	violet
Thymol blue	orange	Bromocresol green	yellow
Methyl yellow	red		

What is the approximate pH of the solution?

2. An aqueous solution of NH_3 has a pH of 11.6. The ammonia (NH_3) molecule is the conjugate base of the NH_4^+ (called "ammonium") ion. Write the net ionic equation for the reaction that makes an aqueous solution of NH_3 basic. (Eq. 5.)

3. The pH of a 0.10 M HOBr solution is 4.8.

- a. What is $[\text{H}^+]$ in that solution?

_____ M

- b. What is $[\text{OBr}^-]$? What is $[\text{HOBr}]$? (Where do the H^+ and OBr^- ions come from?)

_____ M; _____ M

- c. What is the value of K_a for HOBr? What is the value of $\text{p}K_a$?

4. Formic acid, HFor, has a K_a value equal to about 1.8×10^{-4} . A student is asked to prepare a buffer having a pH of 3.90 from a solution of formic acid and a solution of sodium formate having the same molarity. How many milliliters of the NaFor solution should she add to 20. mL of the HFor solution to make the buffer? (See discussion of buffers.)

_____ mL

5. How many mL of 0.10 M NaOH should the student add to 20 mL 0.10 M HFor if she wished to prepare a buffer with a pH of 3.90, the same as in Problem 4?

_____ mL

Advanced Study Assignment Experiment 23

Marcel A. Müller



Advanced Study Assignment Experiment 23:

Research and Technology Objectives and Plans Summary, 1990 **NASA Technical Memorandum**, 1990

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Table of Contents Advanced Study Assignment Experiment 23

1. Understanding the eBook Advanced Study Assignment Experiment 23
 - The Rise of Digital Reading Advanced Study Assignment Experiment 23
 - Advantages of eBooks Over Traditional Books
2. Identifying Advanced Study Assignment Experiment 23
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Advanced Study Assignment Experiment 23
 - User-Friendly Interface
4. Exploring eBook Recommendations from Advanced Study Assignment Experiment 23
 - Personalized Recommendations
 - Advanced Study Assignment Experiment 23 User Reviews and Ratings
 - Advanced Study Assignment Experiment 23 and Bestseller Lists

5. Accessing Advanced Study Assignment Experiment 23 Free and Paid eBooks
 - Advanced Study Assignment Experiment 23 Public Domain eBooks
 - Advanced Study Assignment Experiment 23 eBook Subscription Services
 - Advanced Study Assignment Experiment 23 Budget-Friendly Options
6. Navigating Advanced Study Assignment Experiment 23 eBook Formats
 - ePub, PDF, MOBI, and More
 - Advanced Study Assignment Experiment 23 Compatibility with Devices
 - Advanced Study Assignment Experiment 23 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Advanced Study Assignment Experiment 23
 - Highlighting and Note-Taking Advanced Study Assignment Experiment 23
 - Interactive Elements Advanced Study Assignment Experiment 23
8. Staying Engaged with Advanced Study Assignment Experiment 23
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Advanced Study Assignment Experiment 23
9. Balancing eBooks and Physical Books Advanced Study Assignment Experiment 23
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Advanced Study Assignment Experiment 23
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Advanced Study Assignment Experiment 23
 - Setting Reading Goals Advanced Study Assignment Experiment 23
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Advanced Study Assignment Experiment 23
 - Fact-Checking eBook Content of Advanced Study Assignment Experiment 23
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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