

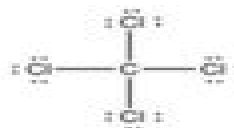
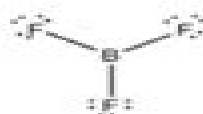
AP Chemistry
Unit 3 Practice Problems
Key

- When two atoms approach each other, the electrons and nucleus of one atom interact with the e⁻ and nucleus of the other. We calculate the effect of these interactions on the energies of the electrons. As atoms approach the overall energy of the system decreases (or becomes more negative) until the atoms are bonded.
 - If the potential energy decreases and bonds form, then energy is released. If atoms are separated and bonds are broken the potential energy increases and energy is released.
 - As the atoms approach and the optimal separation distance is obtained, their orbitals overlap and a bond is formed. The system is at it's lowest energy. If the atoms are forced to get closer the repulsion from the nuclei causes an increase in potential energy.
 - Bond B shows that the potential energy is lower than bond A, therefore more energy released.
 - Bond A is shorter based on the graph.

2. $2(413) + 839 = 1665 \text{ kJ/mol}$

3. $4(413) + 614 + 436 = 2702 \text{ kJ/mol}$ must be absorbed to break all reactant bonds
 $6(413) + 348 = 2826 \text{ kJ/mol}$ is released when product bonds are formed
 $\Delta H_{\text{rxn}} = 2702 - 2826 = -124 \text{ kJ/mol}$

4. a) nonpolar b) polar c) ionic d) polar e) nonpolar f) polar (technically, but it is often considered



6.

	1	2	13	14	15	16	17	18
a	X^1	X^2	X^3	X^4	X^5	X^6	X^7	X^8
b								
c	1+	2+	3+		3-	2-	1-	
d	forms ionic bonds			g	g	g	g	
What combinations of ions are impossible?							AB & CD	
What would have the greatest lattice energy?							AC	
What would have the least lattice energy?							BD	

7.

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