

Name: Key

Class: \_\_\_\_\_

Date: \_\_\_\_\_

Stoichiometry Problem Set A

**Directions:** Solve each of the following problems. Show work using the factor label method or the irf box. Include correct units and significant figures with each answer. Clearly indicate final answer. Be sure to start with a balanced equation

**Part 1: Straight Stoichiometry Problems**

**Target Goal:** To properly calculate the amount of one substance equal to a given amount of another substance within a chemical reaction using the factor-label method.

- 1.) For the reaction  $2\text{HNO}_3 + 1\text{Mg(OH)}_2 \rightarrow 1\text{Mg(NO}_3)_2 + 2\text{H}_2\text{O}$ , how many grams of magnesium <sup>B</sup> nitrate are produced from 8.00 mol of nitric acid ( $\text{HNO}_3$ )? A

$$\text{MM} = 148.3\text{g}$$

$$\frac{8.00\text{mol HNO}_3}{1} \times \frac{1\text{mol Mg(NO}_3)_2}{2\text{mol HNO}_3} \times \frac{148.3\text{g Mg(NO}_3)_2}{1\text{mol Mg(NO}_3)_2} = 592.8\text{g Mg(NO}_3)_2$$

- 2.) For the reaction  $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$ , how many moles of potassium <sup>B</sup> chlorate are required to produce 160. g of oxygen? A

$$\text{MM} = 32.0\text{g}$$

$$\frac{160.\text{g O}_2}{1} \times \frac{1\text{mol O}_2}{32.0\text{g O}_2} \times \frac{2\text{mol KClO}_3}{3\text{mol O}_2} = 3.33\text{ mole KClO}_3$$

- 3.) For the reaction  $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$ , how many grams of water are produced from 5.22 liters of oxygen gas? <sup>B</sup> A

$$\text{MM} = 18.0\text{g}$$

$$\frac{5.22\text{L O}_2}{1} \times \frac{1\text{mol O}_2}{22.4\text{L O}_2} \times \frac{2\text{mol H}_2\text{O}}{1\text{mol O}_2} \times \frac{18.0\text{g H}_2\text{O}}{1\text{mol H}_2\text{O}} = 8.39\text{g H}_2\text{O}$$

- 4.) In the reaction  $2\text{Na} + 2\text{H}_2\text{O} \rightarrow 2\text{NaOH} + \text{H}_2$ , how many grams of hydrogen <sup>B</sup> are produced if 1.33 grams of sodium react completely? A

$$\text{MM} = 23.0\text{g}$$

$$\text{MM} = 2.0\text{g}$$

$$\frac{1.33\text{g Na}}{1} \times \frac{1\text{mol Na}}{23.0\text{g Na}} \times \frac{1\text{mole H}_2}{2\text{mole Na}} \times \frac{2.0\text{g H}_2}{1\text{mole H}_2} = 0.0578\text{g H}_2$$

- 5.) In the reaction  $1\text{CO}_2 + 1\text{H}_2\text{O} \rightarrow 1\text{H}_2\text{CO}_3$ , how many moles of carbonic <sup>B</sup> acid can be produced from 97.2 g of  $\text{CO}_2$ ? A

$$\text{MM} = 44.0\text{g}$$

$$\frac{97.2\text{g CO}_2}{1} \times \frac{1\text{mole CO}_2}{44.0\text{g CO}_2} \times \frac{1\text{mole H}_2\text{CO}_3}{1\text{mole CO}_2} = 2.21\text{ moles H}_2\text{CO}_3$$

**Did you achieve our target goal?**

**Target Goal:** To properly calculate the amount of one substance equal to a given amount of another substance within a chemical reaction using the factor-label method. Circle your answer.

1  
Feeling VERY confused

2  
Not sure that I get this

3  
I think I am catching on

4  
Doing great! Bring it on!

# 12 Stoichiometry Practice Problems Answers Key

**Siddappa N.Byrareddy**



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**Holt Chemistry** Ralph Thomas Myers, 2004. *Chemistry: The Central Science* Theodore L. Brown, H. Eugene LeMay Jr., Bruce E. Bursten, Catherine Murphy, Patrick Woodward, Steven Langford, Dariusz Sagatys, Adrian George, 2013-10-04. If you think you know the Brown LeMay Bursten Chemistry text, think again. In response to market request, we have created the third Australian edition of the US bestseller Chemistry: The Central Science. An extensive revision has taken this text to new heights. Triple-checked for scientific accuracy and consistency, this edition is a more seamless and cohesive product yet retains the clarity, innovative pedagogy, functional problem-solving, and visuals of the previous version. All artwork and images are now consistent in quality across the entire text. And with a more traditional and logical organisation of the Organic Chemistry content, this comprehensive text is the source of all the information and practice problems students are likely to need for conceptual understanding, development of problem-solving skills, reference, and test preparation.

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**The Software Encyclopedia 2000** Bowker Editorial Staff, 2000-05. *Cambridge IGCSE TM Chemistry Student's Book (Collins Cambridge IGCSE TM)* Chris Sunley, 2021-04-22. Collins IGCSE Chemistry provides complete coverage of the latest Cambridge IGCSE syllabus for Chemistry and is packed full of questions in depth, content, practical, investigative skills, features, and more.

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## **Table of Contents 12 Stoichiometry Practice Problems Answers Key**

1. Understanding the eBook 12 Stoichiometry Practice Problems Answers Key
  - The Rise of Digital Reading 12 Stoichiometry Practice Problems Answers Key
  - Advantages of eBooks Over Traditional Books
2. Identifying 12 Stoichiometry Practice Problems Answers Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an 12 Stoichiometry Practice Problems Answers Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from 12 Stoichiometry Practice Problems Answers Key
  - Personalized Recommendations
  - 12 Stoichiometry Practice Problems Answers Key User Reviews and Ratings
  - 12 Stoichiometry Practice Problems Answers Key and Bestseller Lists
5. Accessing 12 Stoichiometry Practice Problems Answers Key Free and Paid eBooks
  - 12 Stoichiometry Practice Problems Answers Key Public Domain eBooks
  - 12 Stoichiometry Practice Problems Answers Key eBook Subscription Services
  - 12 Stoichiometry Practice Problems Answers Key Budget-Friendly Options
6. Navigating 12 Stoichiometry Practice Problems Answers Key eBook Formats

- ePub, PDF, MOBI, and More
- 12 Stoichiometry Practice Problems Answers Key Compatibility with Devices
- 12 Stoichiometry Practice Problems Answers Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of 12 Stoichiometry Practice Problems Answers Key
  - Highlighting and Note-Taking 12 Stoichiometry Practice Problems Answers Key
  - Interactive Elements 12 Stoichiometry Practice Problems Answers Key
- 8. Staying Engaged with 12 Stoichiometry Practice Problems Answers Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers 12 Stoichiometry Practice Problems Answers Key
- 9. Balancing eBooks and Physical Books 12 Stoichiometry Practice Problems Answers Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection 12 Stoichiometry Practice Problems Answers Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine 12 Stoichiometry Practice Problems Answers Key
  - Setting Reading Goals 12 Stoichiometry Practice Problems Answers Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of 12 Stoichiometry Practice Problems Answers Key
  - Fact-Checking eBook Content of 12 Stoichiometry Practice Problems Answers Key
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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