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AP CHEMISTRY MULTIPLE CHOICE QUESTIONS

Unit 8 – Introduction to Equilibrium



The equilibrium constant for the gas phase reaction above is 95 at 25°C. What is the value of the equilibrium constant for the following reaction at 25°C?



Reversed
+ Doubled

(A) 95²

(B) 95

(C) (95)^{1/2}(D) $\frac{1}{95}$ (E) $\frac{1}{(95)^2}$ 

$C_2H_4(g)$ is reduced by $H_2(g)$ in the presence of a solid platinum catalyst, as represented by the equation above. Factors that could affect the rate of the reaction include which of the following?

- I. Changes in the partial pressure of $H_2(g)$
- II. Changes in the particle size of the platinum catalyst
- III. Changes in the temperature of the reaction system

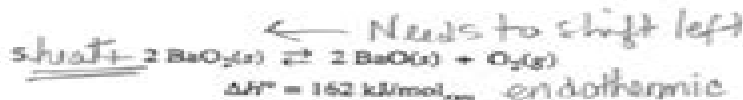
(A) III only

(B) I and II only

(C) I and III only

(D) II and III only

(E) I, II, and III



A sealed rigid vessel contains $BaO_2(s)$ in equilibrium with $BaO(s)$ and $O_2(g)$ as represented by the equation above. Which of the following changes will increase the amount of $BaO_2(s)$ in the vessel?

(A) Removing a small amount of $O_2(g)$ $R \times$ (B) Removing a small amount of $BaO(s)$ \times (C) Adding He gas to the vessel \times (D) Lowering the temperature \checkmark

2. For which of the following reaction mixtures at equilibrium would reducing the volume of the container at constant temperature cause the equilibrium to shift toward the products?

increase P°



shift towards fewer
+ make gas



When 4.00 mol each of $X(g)$ and $Y(g)$ are placed in a 1.00 L vessel and allowed to react at constant temperature according to the equation above, 6.00 mol of $Z(g)$ is produced. What is the value of the equilibrium constant, K_c ?

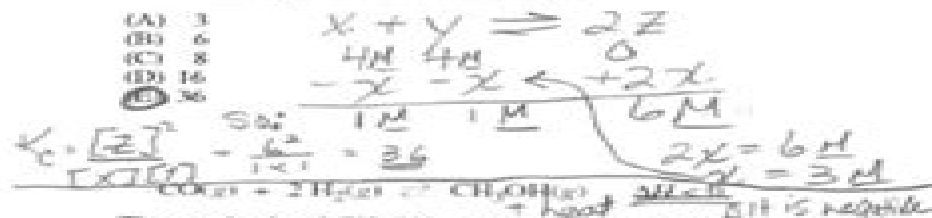
(A) 3

(B) 6

(C) 8

(D) 16

(E) 36



6. The synthesis of $CH_3OH(l)$ from $CO_2(g)$ and $H_2(g)$ is represented by the equation above. The value of K_c for the reaction at 483 K is 14.5.

increase heat shifts left, decrease K_c

Which of the following explains the effect on the equilibrium constant, K_c , when the temperature of the reaction system is increased to 650 K?

(A) K_c will increase because the activation energy of the forward reaction increases more than that of the reverse reaction.(B) K_c will increase because there are more reactant molecules than product molecules.(C) K_c will decrease because the reaction is exothermic.(D) K_c is constant and will not change.

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YCT Expert Team



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story of molecular ionization and its phenomenal evolution that makes mass spectrometry the powerful method it is today Chapters 1 and 2 cover fundamentals and various issues that are common to all ionization e g accurate mass isotope clusters and derivatization Chapters 3 9 acknowledge that some ionization methods are appropriate for gas phase molecules and others for molecules that are in the solid or liquid states Chapters 3 6 cover gas phase molecules dividing the subject into 1 ionization of gas phase molecules by particles e g EI 2 ionization by photons 3 ionization by ion molecule and molecule molecule reactions e g APCI and DART and ionization in Strong electric fields i e Electrohydrodynamic and Field Ionization Desorption Ionization in a Strong Electric Field illustrates the transition to ionization of molecules in the solid or liquid states covered in Chapters 7 9 1 spray methods for ionization e g electrospray 2 desorption ionization by particle bombardment e g FAB and 3 desorption by photons e g MALDI Electrospray and MALDI also lead to applications in biophysical chemistry the theme of Chapter 10 Chapter 11 reconsiders ionization from the view of choosing an ionization method The range of subjects is from ionization of organic and biomolecules to the study of microorganisms **Subject Guide to Books in Print** ,1997

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