

Reference Tables for Physical Setting/EARTH SCIENCE

Radiometric Decay Data

RADIOACTIVE ISOTOPE	DISINTEGRATION	HALF-LIFE (years)
Carbon-14	$^{14}_6\text{C} \rightarrow ^{14}_7\text{N}$	5.7×10^3
Potassium-40	$^{40}_{19}\text{K} \rightarrow ^{40}_{18}\text{Ar}$ $^{40}_{19}\text{K} \rightarrow ^{40}_{20}\text{Ca}$	1.2×10^9
Uranium-238	$^{238}_{92}\text{U} \rightarrow ^{206}_{82}\text{Pb}$	4.5×10^9
Rubidium-87	$^{87}_{37}\text{Rb} \rightarrow ^{87}_{38}\text{Sr}$	4.8×10^{10}

Equations

Acceleration = distance between two lengths of equal size

Gradient = change in field value
distance

Rate of change = change in value
time

Density = mass
volume

Specific Heats of Common Materials

MATERIAL	SPECIFIC HEAT (calories/gm $^{\circ}\text{C}$)
Liquid water	1.00
Boiling water (gas)	0.48
Water vapor	0.48
Dry air	0.17
Granite	0.79
Iron	0.45
Copper	0.39
Lead	0.13

Properties of Water

Heat energy gained during melting	336 J/g
Heat energy released during freezing	336 J/g
Heat energy gained during vaporization	2260 J/g
Heat energy released during condensation	2260 J/g
Density at 4°C	1.00 g/cm ³

Average Chemical Composition of Earth's Crust, Hydrosphere, and Biosphere

ELEMENT (symbol)	CRUST		HYDROSPHERE	
	Percent by mass	Percent by volume	Percent by volume	Percent by volume
Oxygen (O)	46.60	44.54	85.00	85.00
Silicon (Si)	28.20	21.00		
Aluminum (Al)	8.25	8.40		
Iron (Fe)	5.65	5.40		
Calcium (Ca)	4.14	3.56		
Sodium (Na)	2.80	3.10		
Magnesium (Mg)	2.30	2.50		
Potassium (K)	2.58	1.60		
Nitrogen (N)				74.5
Hydrogen (H)			14.0	1.0
Other	0.91	0.97	0.0	0.0

2011 EDITION

This edition of the State Science Assessment Tables should be used in the assessment beginning in the 2011-12 school year. The first examination that includes these tables will be used in the January 2012 Regents Examination in Physical Setting/Earth Science.

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1972 Earth Science Reference Tables

K. M. Lindsay



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