

**Coulomb's Law**

1. The diagram is of a hydrogen atom.
- Label the proton in the nucleus with a  $+$  sign and the orbital electron with a  $-$  sign.
  - The electrical interaction between the nucleus and the orbital electron is a force of \_\_\_\_\_ (attractive) (repulsive).
  - According to Coulomb's law,

$$F = k \frac{q_1 q_2}{r^2}$$

if the charge of either the nucleus or the orbital electron were greater, the force between the nucleus and the electron would be \_\_\_\_\_ (less) (greater) (less).

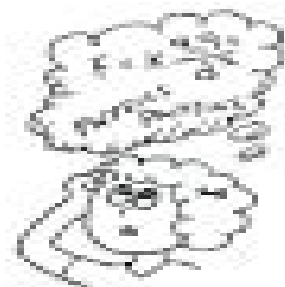
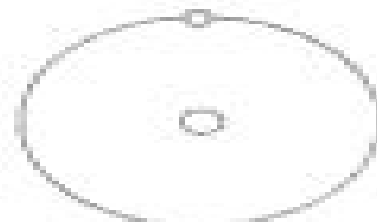
\_\_\_\_\_ (less) (greater) (less).

and if the distance between the nucleus and electron were greater the force would be \_\_\_\_\_ (less) (greater) (less).

If the distance between the nucleus and electron were doubled, the force would be \_\_\_\_\_ (1/4 as much) (1/2 as much) (two times as much) (4 times as much).

2. Consider the electric force between a pair of charged particles a certain distance apart. By Coulomb's law:

- If the charge on one of the particles is doubled, the force is \_\_\_\_\_ (unchanged) (halved) (doubled) (quadrupled).
- If instead the charge on both particles is doubled, the force is \_\_\_\_\_ (unchanged) (halved) (doubled) (quadrupled).
- If instead the distance between the particles is halved, the force is \_\_\_\_\_ (unchanged) (halved) (doubled) (quadrupled).
- If the distance is halved, and the charge of both particles is doubled, the force is \_\_\_\_\_ times as great.



# 32 2 Concept Development

**Padhraic Smyth**



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## **Table of Contents 32 2 Concept Development**

1. Understanding the eBook 32 2 Concept Development
  - The Rise of Digital Reading 32 2 Concept Development
  - Advantages of eBooks Over Traditional Books
2. Identifying 32 2 Concept Development
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an 32 2 Concept Development
  - User-Friendly Interface

4. Exploring eBook Recommendations from 32 2 Concept Development
  - Personalized Recommendations
  - 32 2 Concept Development User Reviews and Ratings
  - 32 2 Concept Development and Bestseller Lists
5. Accessing 32 2 Concept Development Free and Paid eBooks
  - 32 2 Concept Development Public Domain eBooks
  - 32 2 Concept Development eBook Subscription Services
  - 32 2 Concept Development Budget-Friendly Options
6. Navigating 32 2 Concept Development eBook Formats
  - ePub, PDF, MOBI, and More
  - 32 2 Concept Development Compatibility with Devices
  - 32 2 Concept Development Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of 32 2 Concept Development
  - Highlighting and Note-Taking 32 2 Concept Development
  - Interactive Elements 32 2 Concept Development
8. Staying Engaged with 32 2 Concept Development
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers 32 2 Concept Development
9. Balancing eBooks and Physical Books 32 2 Concept Development
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection 32 2 Concept Development
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine 32 2 Concept Development
  - Setting Reading Goals 32 2 Concept Development
  - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of 32 2 Concept Development
  - Fact-Checking eBook Content of 32 2 Concept Development
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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