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web apr 22 2020 molar volume of gases 1 counting in moles one mole is defined as 6.02 x 10<sup>23</sup> particles this is the avogadro

number an unchanging constant whether you are counting hydrogen atoms carbon dioxide molecules or carbonate ions a mole is always  $6.023 \times 10^{23}$  but as different particles have different mass their molar mass varies

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web the molar volume has the si unit of cubic metres per mole  $\text{m}^3 \text{mol}^{-1}$  although it is more typical to use the units cubic decimetres per mole  $\text{dm}^3 \text{mol}^{-1}$  for gases and cubic centimetres per mole  $\text{cm}^3 \text{mol}^{-1}$  for liquids and solids

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web zero degrees celsius or 273 k is defined as standard temperature and one atmosphere or 760 mm hg is defined as standard pressure the volume occupied by one mole of a gas is called molar volume therefore the molar volume of a gas at standard temperature and pressure stp is 22.4 l

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web at standard temperature and pressure stp the molar volume  $V_m$  is the volume occupied by one mole of a chemical element or a chemical compound it can be calculated by dividing the molar mass  $M$  by mass density  $\rho$  molar gas volume is one mole of any gas at a specific temperature and pressure has a fixed volume molar volume formula

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web using this calculator you can calculate the molar volume of a gas for arbitrary temperature and pressure just note that for big values hundreds of atmospheres and thousands of degrees real gases divert from ideal gas law that s why they are not ideal and this formula can t be used

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web this volume is known as the molar gas volume  $v_m$  the volume of a gas may be calculated from its number of moles using volume of gas moles  $v_m$  rearranging this moles of gas  $\frac{v}{v_m}$

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web jun 8 2023  $v = 1.00 \text{ mol} \times 0.08206 \text{ latm mol}^{-1} \text{K}^{-1} \times 273 \text{K} = 22.4 \text{ l}$  in other words 1 mole of a gas will occupy 22.4 l at stp assuming ideal gas behavior at stp the volume of a gas is only dependent on number of moles of

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