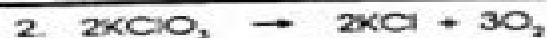


**STOICHIOMETRY:  
MOLE-MOLE PROBLEMS**

How many moles of hydrogen are needed to completely react with two moles of nitrogen?

$$\frac{2 \text{ mol N}_2}{1 \text{ mol N}_2} \times \frac{3 \text{ mol H}_2}{1 \text{ mol N}_2} = 6$$

6 moles H<sub>2</sub>

How many moles of oxygen are produced by the decomposition of six moles of potassium chlorate?

$$6 \text{ moles KClO}_3 \times \frac{3 \text{ mol O}_2}{2 \text{ mol KClO}_3} = \frac{18}{2} = 9$$

9 moles of O<sub>2</sub>

How many moles of hydrogen are produced from the reaction of three moles of zinc with an excess of hydrochloric acid?

$$3 \text{ mol Zn} \times \frac{1 \text{ mol H}_2}{1 \text{ mol Zn}} = 3$$

3 moles of H<sub>2</sub>How many moles of oxygen are necessary to react completely with four moles of propane (C<sub>3</sub>H<sub>8</sub>)?

$$4 \text{ mol C}_3\text{H}_8 \times \frac{5 \text{ mol O}_2}{1 \text{ mol C}_3\text{H}_8} = 20$$

20 mol of O<sub>2</sub>

How many moles of potassium nitrate are produced when two moles of potassium phosphate react with two moles of aluminum nitrate?

$$\frac{2 \text{ mol K}_3\text{PO}_4}{1 \text{ mol K}_3\text{PO}_4} \times \frac{3 \text{ mol KNO}_3}{1 \text{ mol K}_3\text{PO}_4} = 6$$

6 moles of KNO<sub>3</sub>

# Answer Key Stoichiometry Packet

**Julie C. Gilbert**



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reacting volumes of gases with a short cut method A short cut method for solving reacting quantities involving masses and volumes of gases is also given Chapter 3 describes the calculations involved in the practical determination of molecular and empirical formulas A clear meaning of percentage composition of mass is provided and used to solve problems in a step by step manner In chapter 4 we discuss percentage yield and purity A number of examples are given to illustrate how formulas of yield and purity are used in various circumstances A student will find these examples helpful in relating different formulas for percentage purity The last chapter introduces a graphical method for reacting quantities The method may provide a new way of looking at chemical reactions Examples are given to illustrate the method including how it can be used to determine limiting reagents It is hoped that the book will provide all the necessary knowledge and skills to students studying an introductory chemistry course Teachers may also find this book a good resource for their lessons in stoichiometry

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