



that consists of the basic, the components that are present in all matter and cannot be split. In simple terms, that form part of all things around us.

### What is an element?

At  $0^\circ\text{C}$ ,  $100\text{ g}$  of  $\text{H}_2\text{O}$

### What is an atom?

If you look at an element, you cannot see it. You can only see the atoms that make it up. After all, an element is just one type of atom.

At  $0^\circ\text{C}$ ,  $100\text{ g}$  of  $\text{H}_2\text{O}$



At  $0^\circ\text{C}$ ,  $100\text{ g}$  of  $\text{H}_2\text{O}$  consists of  $100\text{ g}$  of water. This means that there are  $100\text{ g}$  of water molecules in the water. These molecules consist of  $100\text{ g}$  of hydrogen and  $100\text{ g}$  of oxygen. This means that there are  $100\text{ g}$  of hydrogen atoms and  $100\text{ g}$  of oxygen atoms in the water.

- It contains two  $\text{H}$  atoms and a  $\text{O}$  atom, what is it called?
- Water is formed in the chemical equation  $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$ . Which side of the equation has the reactants, the left side or the right side?
- Which side of the equation has the products, the left side or the right side?
- What is the most stable electron configuration for an atom?
  - Energy level has an equal number of electrons.
  - The highest energy level has a full set of electrons.
  - There is an even number of electrons.
- Which one of the following sets has an even number of valence electrons?
  - Sharing electrons with another atom.
  - Transferring electrons to another atom.
  - Adding energy to an existing atom.
- Which type of bond forces atoms to move closer together?
  - Ionic bond.
  - Covalent bond.
  - Hydrogen bond.
- When two atoms have an unequal pull on the electrons, they share the bond. This is called a polar bond.
- Parts of the following bond types are the bonds?
  - Ionic bonds.
  - Covalent bonds.
  - Hydrogen bonds.
- Which one of these is shown below?
  - $\text{H}_2\text{O}$
  - $\text{H}_2\text{O}_2$
  - $\text{H}_2\text{O}_3$

Which property of water is not a result of hydrogen bonding?

- Surface tension
- Adhesiveness
- Viscosity
- Adhesion

A water molecule is able to make an  $\text{H}_2\text{O}$  shape being almost  $100^\circ$  because water has hydrogen bonding. This results in a phenomenon called capillary adhesion between the water and glass.

### ADVANCED CONCEPTS

- In the following chemical equation, how many molecules of water are formed in a  $\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$  reaction?
- In the chemical equation in question 13, how many molecules of oxygen are present in a  $\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$  reaction?
- Write how many atoms of each element are present in each side of the following equation and determine which element is present in greatest and the smallest amount.
 
$$\text{NaCl} \rightarrow \text{NaCl} + \text{O}_2$$
- How would you balance the equation in question 13?

# Answer Guide For Compass Learning Physical Science

**Samuel Austin Allibone**

## Answer Guide For Compass Learning Physical Science:

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