

Chemistry Student Edition - Basic Answer Key

Chapter 19: Chemical Equilibrium

The Nature of Chemical Equilibrium

Questions

1. At chemical equilibrium, what is the relationship between the rate of the forward reaction and the rate of the reverse reaction?
2. What does the law of mass action say about any reaction at equilibrium?
3. Why are solids and liquids left out of the equilibrium constant expression?
4. Explain why the equilibrium constant expression as derived from concentrations can be converted to the equilibrium constant expression as derived from partial pressures.
5. Construct balanced chemical equations and write the equilibrium constant expression (either K_c or K_p) for each of the following processes.
 - a) Sodium metal reacts vigorously with oxygen gas to produce sodium oxide.
 - b) Aluminum oxide decomposes to form aluminum metal and oxygen gas.
 - c) Nitrogen gas and oxygen gas react to form nitrogen monoxide gas.
6. Give an example of a chemical equilibrium that exists in our daily lives.
7. Write a K_{eq} expression for each of the following processes.
 - a) Solid calcium carbonate is heated and decomposes into solid calcium oxide and carbon dioxide gas.
 - b) Hydrogen gas (H_2) and solid iodine (I_2) are combined to form hydrogen iodide gas (HI).

Answers

1. At equilibrium, the rate of the forward reaction equals the rate of the reverse reaction.
2. The ratio of product concentrations to reactant concentrations (along with the appropriate exponents) is equal to a constant.
3. Solids and liquids have essentially constant densities, which means that writing them is not very useful.
4. Partial pressures are related to concentrations by a factor of RT .
- 5.



$$K_c = \frac{[Na_{2}O_{(s)}]^2}{[Na_{(s)}]^2 [O_{2(g)}]}$$

- a) heterogeneous equilibria



$$K_c = \frac{[Al_{(s)}]^4 [O_{2(g)}]^3}{[Al_2O_{3(s)}]^2}$$

- b) heterogeneous equilibria

Answer Key Ch 19 Chemistry

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