

Chemistry Student Edition - Basic Answer Key

Chapter 19: Chemical Equilibrium

The Nature of Chemical Equilibrium

Questions

- At chemical equilibrium, what is the relationship between the rate of the forward reaction and the rate of the reverse reaction?
- What does the law of mass action say about any reaction at equilibrium?
- Why are solids and liquids left out of the equilibrium constant expression?
- Explain why the equilibrium constant expression as derived from concentrations can be converted to the equilibrium constant expression as derived from partial pressures.
- Construct balanced chemical equations and write the equilibrium constant expression (either K_c or K_p) for each of the following processes.
 - Sodium metal reacts vigorously with oxygen gas to produce sodium oxide.
 - Aluminum oxide decomposes to form aluminum metal and oxygen gas.
 - Nitrogen gas and oxygen gas react to form nitrogen monoxide gas.
- Give an example of a chemical equilibrium that exists in our daily lives.
- Write a K_{eq} expression for each of the following processes.
 - Solid calcium carbonate is heated and decomposes into solid calcium oxide and carbon dioxide gas.
 - Hydrogen gas (H_2) and solid iodine (I_2) are combined to form hydrogen iodide gas (HI).

Answers

- At equilibrium, the rate of the forward reaction equals the rate of the reverse reaction.
- The ratio of product concentrations to reactant concentrations (along with the appropriate exponents) is equal to a constant.
- Solids and liquids have essentially constant densities, which means that writing them is not very useful.
- Partial pressures are related to concentrations by a factor of RT .
-



$$K_c = \frac{[Na_2O]_2}{[Na]_2 [O]_2}$$

- heterogeneous equilibria



$$K_c = \frac{[Al]_4 [O]_3}{[Al_2O]_2}$$

- heterogeneous equilibria

Answer Key Ch 19 Chemistry

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