

# DISTRIBUTIVE PROPERTY

START

$3 \times (2 + 3)$

$2 \times (3 + 1)$

$5 \times (4 + 2)$

$4 \times (3 + 3)$

$6 \times (2 + 3)$

$9 \times (3 + 4)$



Directions: To play roll a dice, move that number of spaces. Answer the question you land on, if you get it correct you get to stay, if you miss the question move back to the very beginning.

$8 \times (2 + 6)$

$7 \times (5 + 2)$

$4 \times (4 + 5)$

$6 \times (6 + 3)$

$3 \times (5 + 1)$

$4 \times (5 + 2)$

$2 \times (2 + 7)$

$3 \times (6 + 5)$

$5 \times (6 + 2)$

$2 \times (2 + 2)$

$6 \times (4 + 2)$

$7 \times (4 + 4)$

$3 \times (5 + 3)$

$6 \times (6 + 1)$

FINISH

# Distributive Property Games 9th Grade

**Lingsheng Yao**



## **Distributive Property Games 9th Grade:**

**A Comparison of Mathematics Programs for Able Junior High School Students** Miriam L. Goldberg,1966  
**Instructor** ,1977      **Statistics of Land-grant Colleges and Universities** United States. Office of Education,1935  
Teaching Mathematics for the 21st Century Linda Huetinck,Sara N. Munshin,2008 This third edition of Teaching Mathematics for the 21st Century continues to help teachers let the secret out to open up to their students the wonderful discoveries and challenges of the pattern making and problem solving aspects of a fascinating subject mathematics The rationale remains the same to enable prospective and current teachers to access and use tools and strategies to effectively teach mathematics to contemporary students Changing demographics knowledge of how people learn and technology all impact the way we educate our young people This edition incorporates lessons and strategies from programs that have proven success in many types of classrooms Many of these examples help students connect mathematics to real life situations and communicate their understanding of the underlying concepts Although technology is constantly being upgraded ways to increase student motivation through its application remains a goal For example since applets can enhance a lesson whether the teacher uses a computer projector a smart board or has students work individually on computers we have identified several sources of mathematics applets that can be correlated to various lessons Research citations and summaries have been updated to reflect current information on teaching and learning For future teachers      **Teaching Elementary School Mathematics** Klaas Kramer,1975      **Anthropology & Education Quarterly** ,1998      **A Comparison of Mathematics Programs for Able Junior High School Students** Miriam L. Goldberg,1966      **Resources in Education** ,1980  
Research in Education ,1970      **Forest and Stream** ,1895      **The Illustrated London News** ,1870      **Catalog of Copyright Entries. Third Series** Library of Congress. Copyright Office,1978      **The United States Army and Navy Journal and Gazette of the Regular and Volunteer Forces** ,1897      **Chandler's Encyclopedia** William Henry Chandler,1898      **Fanciers' Journal and Poultry Exchange** ,1874      *Golf* ,1890      **The Software Directory for the APPLE Computer** ,1981      **Army-Navy-Air Force Register and Defense Times** ,1936      Chess Life ,2005      **Hockey Field and Lacrosse, Including Net-ball** ,1931

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web other articles where exothermic reaction is discussed alkali metal reactions with nonmetals the reactions are highly exothermic producing up to 235 kcal mole for lithium fluoride the alkali metals react with nonmetals in groups 15 and 16 via and via of the periodic table sulfides can be formed by the direct reaction of the alkali metals with

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web 1 sodium hydroxide solution is poured into a beaker of hydrochloric acid which contains a thermometer showing room temperature 2 the beaker now contains sodium chloride and water and the

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web reacting two solutions eg acid and alkali place the polystyrene cup inside the glass beaker to make it more stable measure an appropriate volume of each liquid eg 25 cm<sup>3</sup> place one of the

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web for reactions involving acetic acid or ammonia the measured enthalpy change of neutralization is a few kJ less exothermic than with strong acids and bases for example one source which gives the enthalpy change of neutralization of sodium hydroxide solution with HCl as 57.9 kJ mol<sup>-1</sup>

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web indicators are used to determine whether a solution is acidic or alkaline acids react with metals bases and carbonates to



produce salts neutralisation is the reaction between an acid and a base

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web sep 20 2014 i will illustrate this with some examples when an acid is neutralised by an alkali the reaction is exothermic

eg 1  $\text{HCl(aq)} + \text{NaOH(aq)} \rightarrow \text{NaCl(aq)} + \text{H}_2\text{O(l)}$  for which  $\Delta H = -57\text{ kJ mol}^{-1}$  eg 2  $\text{HNO}_3\text{(aq)} + \text{KOH(aq)} \rightarrow \text{KNO}_3\text{(aq)} + \text{H}_2\text{O(l)}$  for which  $\Delta H = -57\text{ kJ mol}^{-1}$

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web the enthalpy change of neutralisation is defined as the enthalpy change when solutions of an acid and alkali react together to produce one mole of water for example  $\text{HCl(aq)} + \text{NaOH(aq)} \rightarrow \text{NaCl(aq)} + \text{H}_2\text{O(l)}$  you can see that one mole of water is produced when one mole of HCl or one mole of NaOH reacts