

AP Multiple Choice Questions.

1989

27.

- I. Difference in temperature between freezing point of solvent and freezing point of solution
- II. Molal freezing point depression constant, K_f , for solvent

In addition to the information above, which of the following gives the minimum data required to determine the molecular mass of a nonionic substance by the freezing point depression technique?

- a. No further information is necessary.
- b. Mass of solute
- c. Mass of solute and mass of solvent
- d. Mass of solute and volume of solvent
- e. Mass of solute, mass of solvent, and vapor pressure of solvent

71. A solution of toluene (molecular weight 92.1) in benzene (molecular weight 78.1) is prepared. The mole fraction of toluene in the solution is 0.100. What is the molality of the solution?

- a. 0.100 m
- b. 0.703 m
- c. 0.921 m
- d. 1.28 m
- e. 1.42 m

1994

14. Which of the following is lower for a 1.0-molar aqueous solution of any solute than it is for pure water?

- a. pH
- b. Vapor pressure
- c. Freezing point
- d. Electrical conductivity
- e. Absorption of visible light

44. Which of the following solutions has the lowest freezing point?

2002

26. Approximately what mass of $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$ (250 g/mol) is required to prepare 250 mL of 0.10 M copper(II) sulfate solution?

- (A) 4.0 g
- (B) 6.2 g
- (C) 34 g
- (D) 85 g
- (E) 140 g

38. A 0.10 M aqueous solution of sodium sulfate, Na_2SO_4 , is a better conductor of electricity than a 0.10 M aqueous solution of sodium chloride, NaCl . Which of the following best explains this observation?

- a. Na_2SO_4 is more soluble in water than NaCl is.
- b. Na_2SO_4 has a higher molar mass than NaCl has.
- c. To prepare a given volume of 0.10 M solution, the mass of Na_2SO_4 needed is more than twice the mass of NaCl needed.
- d. More moles of ions are present in a given volume of 0.10 M Na_2SO_4 than in the same volume of 0.10 M NaCl .
- e. The degree of dissociation of Na_2SO_4 in solution is significantly greater than that of NaCl .

45. What is the mole fraction of ethanol, $\text{C}_2\text{H}_5\text{OH}$, in an aqueous solution that is 46 percent ethanol by mass? (The molar mass of $\text{C}_2\text{H}_5\text{OH}$ is 46 g; the molar mass of H_2O is 18 g.)

- a. 0.25

Ap Chemistry Multiple Choice Questions 2014

Julia Schneider

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1. Understanding the eBook Ap Chemistry Multiple Choice Questions 2014
 - The Rise of Digital Reading Ap Chemistry Multiple Choice Questions 2014
 - Advantages of eBooks Over Traditional Books
2. Identifying Ap Chemistry Multiple Choice Questions 2014
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ap Chemistry Multiple Choice Questions 2014
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ap Chemistry Multiple Choice Questions 2014

- Personalized Recommendations
 - Ap Chemistry Multiple Choice Questions 2014 User Reviews and Ratings
 - Ap Chemistry Multiple Choice Questions 2014 and Bestseller Lists
5. Accessing Ap Chemistry Multiple Choice Questions 2014 Free and Paid eBooks
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 - Ap Chemistry Multiple Choice Questions 2014 Budget-Friendly Options
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 - Ap Chemistry Multiple Choice Questions 2014 Compatibility with Devices
 - Ap Chemistry Multiple Choice Questions 2014 Enhanced eBook Features
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 - Highlighting and Note-Taking Ap Chemistry Multiple Choice Questions 2014
 - Interactive Elements Ap Chemistry Multiple Choice Questions 2014
8. Staying Engaged with Ap Chemistry Multiple Choice Questions 2014
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ap Chemistry Multiple Choice Questions 2014
9. Balancing eBooks and Physical Books Ap Chemistry Multiple Choice Questions 2014
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ap Chemistry Multiple Choice Questions 2014
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Ap Chemistry Multiple Choice Questions 2014
- Setting Reading Goals Ap Chemistry Multiple Choice Questions 2014
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Ap Chemistry Multiple Choice Questions 2014

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 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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