

AP Chemistry Chapter 4 Practice Exam

D

1. The Claus reactions, shown below, are used to generate elemental sulfur from hydrogen sulfide.



How much sulfur (in grams) is produced from 21.5 grams of O_2 ?

A) 97.0 g



B) 14.4 g

C) 21.5 g

D) 43.1 g

E) none of these

B

2. Iron is produced from its ore by the reactions:



How many moles of $\text{O}_2(g)$ are needed to produce 4.6 moles of $\text{Fe}(s)$?

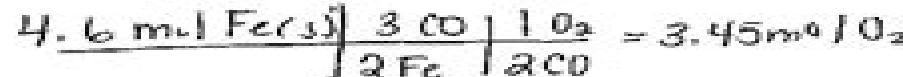
A) 2.3 mole O_2

B) 3.5 mole O_2

C) 4.6 mole O_2

D) 6.9 mole O_2

E) 13.8 mole O_2

**B**

3. Suppose the reaction $\text{Ca}_3(\text{PO}_4)_2 + 3\text{H}_2\text{SO}_4 \rightarrow 3\text{CaSO}_4 + 2\text{H}_3\text{PO}_4$ is carried out starting with 153 g of $\text{Ca}_3(\text{PO}_4)_2$

and 76.8 g of H_2SO_4 . How much phosphoric acid will be produced?

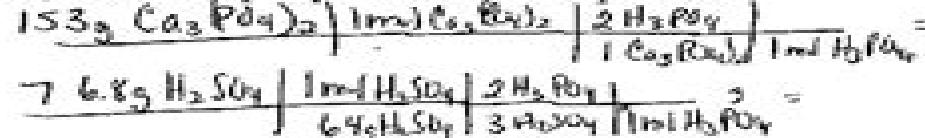
A) 76.7 g

B) 51.1 g

C) 229.8 g

D) 115.1 g

E) 96.7 g

**B**

4. SO_2 reacts with H_2S as follows:



When 7.50 g of H_2S reacts with 12.75 g of SO_2 , which statement applies?

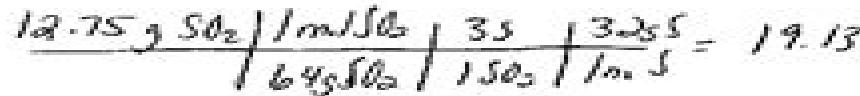
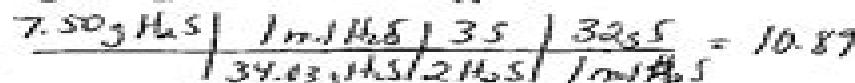
A) 6.38 g of sulfur are formed.

B) 10.6 g of sulfur are formed.

C) 0.0216 moles of H_2S remain.

D) 1.13 g of H_2S remain.

E) SO_2 is the limiting reagent.



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Neil D. Jespersen, Pamela Kerrigan

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