

19.5

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and pOH and concentrations of acids and bases and vice versa.
3. Calculate the pH of weak acids.

Vocabulary

- | | | |
|-------------------|-----------------------------|---------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution |
| 4. autoionization | 5. acid rain | 6. neutral solution |
| | 7. acid solution | 8. pH |

Key Equation

- | | |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

Part A Completion

Write the equation and/or the numerical value(s) of the quantity indicated that are indicated in the problem. Be sure to include units where appropriate.

1. What substance is _____, a weak hydrogen ion donor?
2. _____ is the concentration of H^+ ions present in 100 mL of _____.
3. _____ is the concentration of _____ in _____.
4. _____ is the concentration of _____ in _____.
5. _____ is the concentration of _____ in _____.
6. _____ is the concentration of _____ in _____.
7. _____ is the concentration of _____ in _____.
8. _____ is the concentration of _____ in _____.
9. _____ is the concentration of _____ in _____.
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15. _____ is the concentration of _____ in _____.
16. _____ is the concentration of _____ in _____.
17. _____ is the concentration of _____ in _____.
18. _____ is the concentration of _____ in _____.
19. _____ is the concentration of _____ in _____.
20. _____ is the concentration of _____ in _____.

1. _____

2. _____

3. _____

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11. _____

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17. _____

18. _____

19. _____

20. _____

19 2 Hydrogen Ions And Acidity Work Sheet Answers

Wilhelm Mansfield Clark



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