



# HYDROGEN IONS AND ACIDITY

## Section Summary

### Objectives

- Define hydrogen ion concentration and relate it to the pH of a solution.
- Explain how the concentration of hydrogen ions in a solution is measured.
- Describe strong and weak acids.

### Key Vocabulary

<ul style="list-style-type: none"><li>concentration</li><li>hydrogen ion</li><li>ion</li><li>logarithm</li><li>logarithmic scale</li><li>acid</li><li>strong acid</li><li>weak acid</li><li>base</li><li>strong base</li><li>weak base</li><li>pH</li><li>hydroxide ion</li><li>conjugate acid</li><li>conjugate base</li></ul>	<ul style="list-style-type: none"><li>concentration</li><li>hydrogen ion</li><li>ion</li><li>logarithm</li><li>logarithmic scale</li><li>acid</li><li>strong acid</li><li>weak acid</li><li>base</li><li>strong base</li><li>weak base</li><li>pH</li><li>hydroxide ion</li><li>conjugate acid</li><li>conjugate base</li></ul>	<ul style="list-style-type: none"><li>concentration</li><li>hydrogen ion</li><li>ion</li><li>logarithm</li><li>logarithmic scale</li><li>acid</li><li>strong acid</li><li>weak acid</li><li>base</li><li>strong base</li><li>weak base</li><li>pH</li><li>hydroxide ion</li><li>conjugate acid</li><li>conjugate base</li></ul>
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### Key Equations

- $\text{pH} = -\log_{10}[\text{H}^+]$
- $\text{pH} = \log_{10}[\text{OH}^-]$

$$\text{pH} = -\log_{10}[\text{H}^+]$$

### Practice Calculations

Use the following equations to calculate the pH of a solution with a hydrogen ion concentration of  $1.0 \times 10^{-4} \text{ M}$ .

$\text{pH} = -\log_{10}[\text{H}^+]$        $\text{pH} = -\log_{10}[\text{H}^+]$   
concentration      concentration

Concentration	pH
$1.0 \times 10^{-4} \text{ M}$	4.00
$1.0 \times 10^{-5} \text{ M}$	5.00
$1.0 \times 10^{-6} \text{ M}$	6.00
$1.0 \times 10^{-7} \text{ M}$	7.00
$1.0 \times 10^{-8} \text{ M}$	8.00
$1.0 \times 10^{-9} \text{ M}$	9.00
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**Wilhelm Mansfield Clark**

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