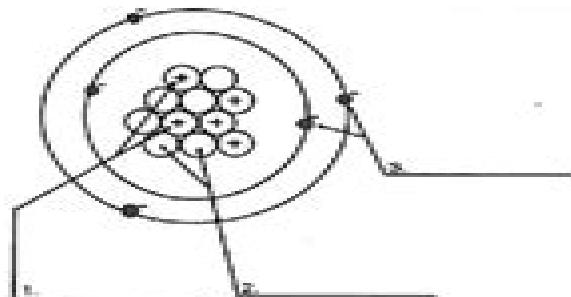


## Atomic Structure Worksheet

Label the parts of an atom on the diagram below.



4. What type of charge does a proton have? + neutral
5. What type of charge does a neutron have? -
6. What type of charge does an electron have? -
7. Which two subatomic particles are located in the nucleus of an atom? p + n
8. What is the charge of an atom? neutral
9. What is the charge of an ion? (Explain your answer fully.)  
+ or - gains or loses e<sup>-</sup>
10. If an atom has 35 protons in the nucleus, how many electrons will it have orbiting the nucleus? 35
11. What is the atomic number of the atom in the diagram above? 5
12. What is the atomic mass/mass number of the atom in the diagram above? 11
13. How many protons are in the nucleus of an atom with an atomic number of 15? 15
14. How many electrons are in the nucleus of an atom with an atomic number of 20? 20
15. How many neutrons are in the nucleus of an atom with an atomic number of 25? (use Periodic Table for mass)  
\_\_\_\_\_
16. What is the mass number of an atom with 3 protons, 4 neutrons, and 3 electrons? 7
17. How many neutrons are in the nucleus of an atom that has an atomic mass of 36 and an atomic number of 25? 11
18. The atomic number tells you the number of protons in one atom of an element. It also tells you the number of electrons in a neutral atom of that element. The atomic number gives the "identity" of an element as well as it's location on the Periodic Table. No two different elements will have the same atomic number.
19. The mass number is used to calculate the number of p, n, e<sup>?</sup> in one atom of an element. In order to calculate the number of neutrons you must subtract the # p from the atomic #.  
(atomic #)

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