

AP\* Chemistry  
**ELECTROCHEMISTRY**

**Terms to Know:**

**Electrochemistry**—the study of the interchange of chemical and electrical energy

**OIL RIG** — oxidation is loss, reduction is gain (of electrons)

**Oxidation** — the loss of electrons, increase in charge

**Reduction** — the gain of electrons, reduction of charge

**Oxidation number** — the assigned charge on an atom

**Oxidizing agent (OA)** — the species that is reduced and thus causes oxidation

**Reducing agent (RA)** — the species that is oxidized and thus causes reduction

**ELECTROCHEMISTRY INVOLVES TWO MAIN TYPES OF PROCESSES:**

- Galvanic (voltaic) cells — which are spontaneous chemical reactions (battery)
- Electrolytic cells — which are non-spontaneous and require external  $e^-$  source (DC power source)
- BOTH of these fit into the category entitled **Electrochemical cells**

**GALVANIC CELLS**

*Parts of the voltaic or galvanic cell:*

- Anode**—the electrode where oxidation occurs. After a period of time, the anode may appear to become smaller as it falls into solution.
- Cathode**— the anode where reduction occurs. After a period of time it may appear larger, due to ions from solution plating onto it.
- Electrodes**— used when a gas is involved OR ion to ion involved such as  $Fe^{3+}$  being reduced to  $Fe^{2+}$  rather than  $Fe^0$ . Made of Pt or graphite.
- Salt bridge** — a device used to maintain electrical neutrality in a galvanic cell. This may be filled with agar which contains a neutral salt or it may be replaced with a porous cup.
- Electron flow** — always from anode to cathode. (through the wire)
- Standard cell notation (line notation)** - anode/solution//cathode solution/ cathode Ex.  $Zn/Zn^{2+}$  (1.0 M) //  $Cu^{2+}$  (1.0 M) / Cu
- Voltmeter** - measures the cell potential ( $emf$ ). Usually is measured in volts.

Balance this re-dox reaction:



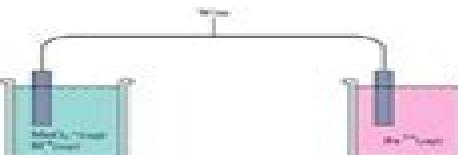
RED:

OA:

OX:

RA

Overall rxn:



# Ap Chemistry Electrochemistry Answers

**Mary Millhollon, Richard H. Langley**

## Ap Chemistry Electrochemistry Answers:

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