

Section 5.3 Electron Configuration

In your textbook, read about ground-state electron configurations.

Use each of the terms below just once to complete the passage.

Aufbau principle lowest	electron configuration	ground-state electron configuration	Hund's rule stable
	Pauli exclusion principle	spins	

The arrangement of electrons in an atom is called the atom's

(1) _____. Electrons in an atom tend to assume the arrangement that gives the atom the (2) _____ possible energy. This arrangement of electrons is the most (3) _____ arrangement and is called the atom's (4) _____.

Three rules define how electrons can be arranged in an atom's orbitals. The

(5) _____ states that each electron occupies the lowest energy orbital available. The (6) _____ states that a maximum of two electrons may occupy a single atomic orbital, but only if the electrons have opposite (7) _____. (8) _____ states that single electrons with the same spin must occupy each equal-energy orbital before additional electrons with opposite spins occupy the same orbitals.

Complete the following table.

Element	Atomic Number	Orbitals					Electron Configuration
		1s	2s	2p _x	2p _y	2p _z	
9. Helium	2						1s ²
10.	7						
11. Neon	10	↑↓	↑↓	↑↓	↑↓	↑↓	

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