

(d) The halogen fluorine is too reactive to use in a school or college laboratory. Fluorine is a powerful oxidising agent. It will react with water as shown below.

(i) Complete and balance the equation for the reaction of fluorine with water.

$$\text{F}_2(\text{g}) + \text{H}_2\text{O}(\text{l}) \rightarrow \text{.....} + \text{O}_2(\text{g}) \quad [1]$$

(ii) Using oxidation numbers, show what has been oxidised and what has been reduced in this reaction.

[21]

(e) Fluorine will react violently with gallium to produce gallium fluoride.

Mendeleev originally called gallium 'eka-aluminium' as he predicted that gallium would have similar properties to aluminium.

(b) Complete the electron structure of the gallium atom.

$1s^2$ [1]

(b) Use Mendeleev's prediction to suggest the empirical formula of gallium fluoride.

11

[Total: 19]

END OF QUESTION PAPER



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