

Multiple Choices: Entropy

1. In which of the following does the entropy decrease?

- A) $\text{NaCl(s)} \rightarrow \text{Na}^+(\text{aq}) + \text{Cl}^-(\text{aq})$
 B) $4\text{NO(g)} + 6\text{H}_2\text{O(g)} \rightarrow 4\text{NH}_3(\text{g}) + 5\text{O}_2(\text{g})$
 C) $2\text{NaHCO}_3(\text{s}) \rightarrow \text{Na}_2\text{CO}_3(\text{s}) + \text{CO}_2(\text{g}) + \text{H}_2\text{O(g)}$
 D) $\text{CaCO}_3(\text{s}) + 2\text{HCl(aq)} \rightarrow \text{CaCl}_2(\text{aq}) + \text{CO}_2(\text{g}) + \text{H}_2\text{O(l)}$

ANS: B

1. The normal boiling point of methanol is 64.7°C and the molar enthalpy of vaporization is 71.6 kJ/mol . The value of ΔS when 2.15 mol of $\text{CH}_3\text{OH(l)}$ vaporizes at 64.7°C is _____ J/K .

- A) 0.457 B) 521×10^3
 C) 457 D) 2.39×10^3
 E) 2.39

ANS: C

1. Which one of the following reactions has a positive entropy change?

- A) $\text{H}_2\text{O(g)} \rightarrow \text{H}_2\text{O(l)}$
 B) $\text{BF}_3(\text{g}) + \text{NH}_3(\text{g}) \rightarrow \text{F}_3\text{BNH}_3(\text{s})$
 C) $2\text{SO}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2\text{SO}_3(\text{g})$
 D) $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightarrow 2\text{NH}_3(\text{g})$
 E) $2\text{NH}_4\text{NO}_3(\text{s}) \rightarrow 2\text{N}_2(\text{g}) + 4\text{H}_2\text{O(g)} + \text{O}_2(\text{g})$

ANS: E

1. For a certain process at 127°C , $\Delta G = -16.20\text{ kJ}$ and $\Delta H = -17.0\text{ kJ}$. What is the entropy change for this process at this temperature? Express your answer in the form, $\Delta S = \frac{\text{J/K}}$.

- A) -6.3 J/K B) $+6.3\text{ J/K}$ C) -2.0 J/K
 D) $+2.0\text{ J/K}$ E) -0.1 J/K

ANS: C

1. Evaluate ΔS° for the reaction below at 25°C .



	$\text{CH}_4(\text{g})$	$+ 2\text{Cl}_2(\text{g})$	\rightarrow	$\text{CCl}_4(\text{l})$	$+ 2\text{H}_2(\text{g})$
$\Delta G^\circ_f\text{ (kJ/mol)}$	-50.8	0		-63.9	0
$\Delta G^\circ_f\text{ (kJ/mol)}$	-16.7	0		-67.7	0

- A) -360 J/K B) -66.9 J/K C) -155 J/K
 D) -467 J/K E) -387 J/K

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Mary Millhollon, Richard H. Langley



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Table of Contents Ap Chemistry Multiple Choice Questions And Answers

1. Understanding the eBook Ap Chemistry Multiple Choice Questions And Answers
 - The Rise of Digital Reading Ap Chemistry Multiple Choice Questions And Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Ap Chemistry Multiple Choice Questions And Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ap Chemistry Multiple Choice Questions And Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ap Chemistry Multiple Choice Questions And Answers
 - Personalized Recommendations
 - Ap Chemistry Multiple Choice Questions And Answers User Reviews and Ratings
 - Ap Chemistry Multiple Choice Questions And Answers and Bestseller Lists
5. Accessing Ap Chemistry Multiple Choice Questions And Answers Free and Paid eBooks
 - Ap Chemistry Multiple Choice Questions And Answers Public Domain eBooks
 - Ap Chemistry Multiple Choice Questions And Answers eBook Subscription Services
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6. Navigating Ap Chemistry Multiple Choice Questions And Answers eBook Formats

- ePub, PDF, MOBI, and More
 - Ap Chemistry Multiple Choice Questions And Answers Compatibility with Devices
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7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ap Chemistry Multiple Choice Questions And Answers
 - Highlighting and Note-Taking Ap Chemistry Multiple Choice Questions And Answers
 - Interactive Elements Ap Chemistry Multiple Choice Questions And Answers
 8. Staying Engaged with Ap Chemistry Multiple Choice Questions And Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ap Chemistry Multiple Choice Questions And Answers
 9. Balancing eBooks and Physical Books Ap Chemistry Multiple Choice Questions And Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ap Chemistry Multiple Choice Questions And Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Ap Chemistry Multiple Choice Questions And Answers
 - Setting Reading Goals Ap Chemistry Multiple Choice Questions And Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Ap Chemistry Multiple Choice Questions And Answers
 - Fact-Checking eBook Content of Ap Chemistry Multiple Choice Questions And Answers
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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