

## Multiple Choices: Entropy

1. In which of the following does the entropy decrease?

- A)  $\text{NaCl}_{(s)} \rightarrow \text{Na}^{+}_{(aq)} + \text{Cl}^{-}_{(aq)}$
- B)  $4\text{Na}_{(s)} + \text{S}^{2-}_{(aq)} \rightarrow 4\text{Na}^{+}_{(aq)} + \text{S}^{2-}_{(aq)}$
- C)  $2\text{NaNO}_3_{(s)} \rightarrow \text{Na}_2\text{O}_{(s)} + \text{O}_2(g) + \text{NaCl}_{(s)}$
- D)  $\text{CaCO}_3_{(s)} + 2\text{HCl}_{(aq)} \rightarrow \text{CaCl}_{(aq)} + \text{CO}_2(g) + \text{H}_2\text{O}_{(l)}$

ANS: B

1. The normal boiling point of methanol is  $-91.7^\circ\text{C}$  and the molar enthalpy of vaporization is  $73.0 \text{ kJ/mol}$ . The value of  $\Delta S$  when 2.15 mol of  $\text{CH}_3\text{OH}_{(l)}$  vaporizes at  $-91.7^\circ\text{C}$  is \_\_\_\_\_.

- A) 0.457
- B)  $5.21 \times 10^3$
- C) 457
- D)  $1.39 \times 10^3$
- E) 2.39

ANS: C

1. Which one of the following reactions has a positive entropy change?

- A)  $\text{NaCl}_{(s)} \rightarrow \text{Na}^{+}_{(aq)} + \text{Cl}^{-}_{(aq)}$
- B)  $\text{S}_{(s)} + 2\text{NH}_3_{(g)} \rightarrow \text{NH}_4\text{NH}_3_{(s)}$
- C)  $2\text{H}_2\text{O}_{(g)} + \text{O}_{(g)} \rightarrow 2\text{H}_2\text{O}_{(l)}$
- D)  $\text{Na}_{(s)} + 3\text{H}_{(g)} \rightarrow 2\text{NaH}_{(s)}$
- E)  $2\text{NaH}_2\text{O}_{(g)} \rightarrow 2\text{Na}_{(s)} + 4\text{H}_2\text{O}_{(g)} + \text{O}_{(g)}$

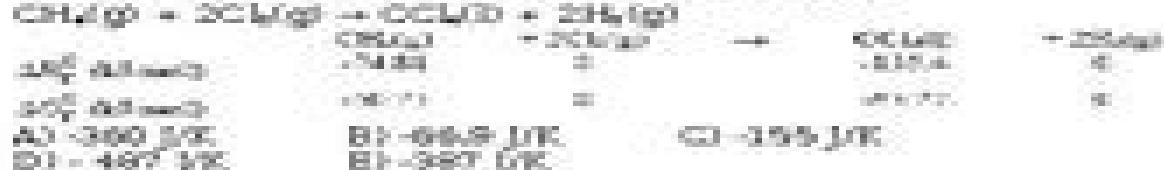
ANS: B

1. For a certain process at  $127^\circ\text{C}$ ,  $\Delta G = -16.00 \text{ kJ}$  and  $\Delta H = -1700 \text{ kJ}$ . What is the entropy change for this process at this temperature? Express your answer in the form,  $\Delta S = \text{_____ J/K}$ .

- A)  $-4.3 \text{ J/K}$
- B)  $-6.3 \text{ J/K}$
- C)  $-2.0 \text{ J/K}$
- D)  $-2.0 \text{ J/K}$
- E)  $-0.1 \text{ J/K}$

ANS: C

1. Evaluate  $\Delta S^\circ$  for the reaction below at  $27^\circ\text{C}$ .



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**Mary Millhollon, Richard H. Langley**

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