

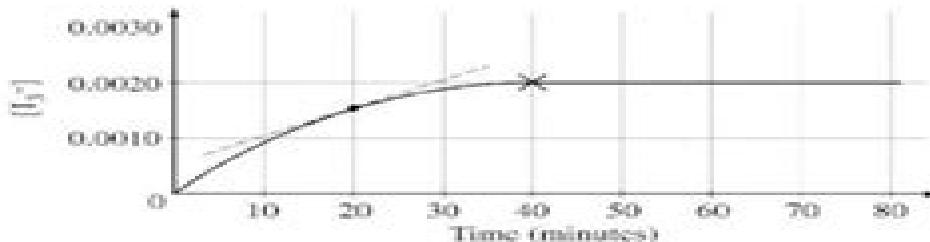
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**Question 6**

(10 points)

(a) The reaction is an oxidation-reduction (redox) reaction. *1 point*  
To go from  $I^-$  to  $I_3^-$ , the iodine has to be oxidized, so credit is earned for the formula of any reasonable oxidizing agent (e.g.,  $F_2$ ,  $Co^{3+}$ ,  $O_2$ ,  $MnO_4^-$ ) *1 point*

(b) (i) One point is earned for an "X" drawn on the curve at a point corresponding to approximately 37–42 minutes (see graph below). This is where the curve levels off. *1 point*  
(ii) The instantaneous rate of the reaction at 20 minutes is the slope of the line tangent to the curve at 20 minutes. (See the graph below.) *1 point*



- No credit is earned if the student confuses "rate" with "rate constant".

(b) Keep the initial concentration of  $S_2O_8^{2-}$  constant, change the initial concentration of  $I^-$  by a known amount, and see what effect this change has on the measured rate of the reaction. This procedure should be repeated, but this time keeping the concentration of  $I^-$  constant and changing the concentration of  $S_2O_8^{2-}$ . *2 points*

Note: Points are earned for a method that involves controlling one reactant concentration and varying the other while comparing the reaction rates.

- No credit is earned if student confuses product and reactant concentrations.

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**L Manion**

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