

19.2

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate K_a and K_b for weak acids and bases, and calculate pK_a and pK_b for weak acids and bases.
2. Calculate K_a and K_b for weak acids and bases, and calculate pK_a and pK_b for weak acids and bases.
3. Calculate the pH of weak acid and base solutions.

Vocabulary

- | | | |
|-----------------|--------------------------------------|-------------------|
| acid | acid dissociation constant (K_a) | acid ionization |
| acidic solution | acid strength | alkaline solution |
| acidic solution | acid strength | alkaline solution |

Key Equations

- | | |
|---|--|
| $K_a = \frac{[H^+][A^-]}{[HA]}$ (Equation 19.1) | $K_b = \frac{[OH^-][HB^+]}{[B]}$ (Equation 19.2) |
|---|--|

Part A Completion

Write the chemical formula for each of the following acids, bases, and salts, showing the correct number of atoms of each element within each molecule or formula.

1. Nitric acid
2. Acetic acid
3. Hydrochloric acid
4. Sulfuric acid
5. Phosphoric acid
6. Nitrous acid
7. Hydrofluoric acid
8. Hydrobromic acid
9. Hydroiodic acid
10. Hydrocyanic acid
11. Hydroperchloric acid
12. Hydroperbromic acid
13. Hydroperiodic acid
14. Hydroperacetic acid
15. Hydroperacetic acid
16. Hydroperacetic acid
17. Hydroperacetic acid
18. Hydroperacetic acid
19. Hydroperacetic acid
20. Hydroperacetic acid

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19 2 Hydrogen Ions And Acidity Section Review

David Kirk



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