

19.2

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate K_a and K_b for weak acids and bases, and vice versa for weak acids and bases.
2. Calculate K_a and K_b for weak acids and bases, and vice versa for weak acids and bases.
3. Calculate the pH of weak acids and bases.

Vocabulary

- acid
- base

- acidic solution
- basic solution

- acidic solution
- basic solution

Key Equations

- 1. $K_a = \frac{[H^+][A^-]}{[HA]}$ (Equation 19.1)

- 2. $K_b = \frac{[OH^-][BH^+]}{[B]}$ (Equation 19.2)

Part A Completion

Write the chemical formula for each of the following acids, bases, and salts, showing the correct number of atoms of each element within each molecule or formula.

1. Hydrochloric acid _____

2. Acetic acid _____

3. Nitric acid _____

4. Sulfuric acid _____

5. Phosphoric acid _____

6. Hydrofluoric acid _____

7. Hydrobromic acid _____

8. Hydroiodic acid _____

9. Hydrocyanic acid _____

10. Hydroperchloric acid _____

11. Hydroperbromic acid _____

12. Hydroperiodic acid _____

13. Hydroperfluoric acid _____

14. Hydroperchloric acid _____

15. Hydroperbromic acid _____

16. Hydroperiodic acid _____

17. Hydroperfluoric acid _____

18. Hydroperchloric acid _____

19. Hydroperbromic acid _____

20. Hydroperiodic acid _____

21. Hydroperfluoric acid _____

22. Hydroperchloric acid _____

19 2 Hydrogen Ions And Acidity Section Review

G Thomas



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