

16. Increasing the temperature generally ...

- decreases the solubility of a solid in a liquid but increases the solubility of a gas in a liquid.
- increases the solubility of a solid in a liquid but decreases the solubility of a gas in a liquid.
- increases the solubility of both a solid and a gas in a liquid.
- decreases the solubility of both a solid and a gas in a liquid.

17. A white precipitate was formed when $\text{BaCl}_{2(\text{aq})}$ was added to an aqueous solution of a salt X. The precipitate dissolved in dilute HCl with rapid effervescence. Salt X is likely to contain

- SO_4^{2-} ions.
- SO_3^{2-} ions.
- NO_3^- ions.
- CO_3^{2-} ions.

18. Before a reaction could take place, there should be

- ionization of reactant particles.
- breakage of bonds of products.
- breakage of bonds of reactants.
- ionization of product particles.

19. Consider the following reaction equation:

$$\text{CaO} + \text{SiO}_2 \rightarrow \text{CaSiO}_3$$

 Silicon(IV) oxide is acting as

- a basic oxide.
- a reducing agent.
- an acidic oxide.
- an oxidizing agent.

20. Which of the following acids would form normal salt only?

- Tetraoxosulphate(VI) acid
- Trioxonitrate(V) acid
- Tetraoxophosphate(V) acid
- Trioxosulphate(IV) acid

21. What is the partial pressure of oxygen at s.t.p. in a gaseous mixture containing 100 cm³ of oxygen and 900 cm³ of nitrogen gas?

- 1.0 atm
- 0.1 atm
- 9.0 atm
- 0.9 atm

22. Graphite is used as a dry lubricant due to the presence of

- mobile electrons.
- free electrons.
- octahedral structures.
- layered structures.

23. Which of the following gases is alkaline?

- NO_2
- CO_2
- NH_3
- N_2O

24. Which of the following statements about an equilibrium system is correct?

- Forward and backward reactions occur at the same rate.
- The concentrations of reactants must equal that of the products.
- The concentrations of reactants and products can be changed by adding a catalyst.
- The concentrations of reactants and products are not affected by a change in temperature.

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